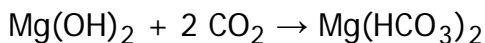


# Magnesium bicarbonate

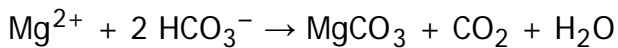
From Wikipedia, the free encyclopedia

**Magnesium bicarbonate** or **magnesium hydrogen carbonate**, Mg(HCO<sub>3</sub>)<sub>2</sub>, is the **bicarbonate** salt of **magnesium**. It can be formed through the reaction of dilute solutions of **carbonic acid** (such as **seltzer water**) and **magnesium hydroxide** (milk of magnesia).

Magnesium bicarbonate exists only in aqueous solution. To produce it, a suspension of magnesium hydroxide is treated with pressurized **carbon dioxide**, producing a solution of magnesium bicarbonate:<sup>[1]</sup>



Drying the resulting solution causes the magnesium bicarbonate to **decompose**, yielding **magnesium carbonate**, carbon dioxide, and water:



## References [edit]

- ↑ Margarete Seeger; Walter Otto; Wilhelm Flick; Friedrich Bickelhaupt; Otto S. Akkerman (2005), "Magnesium Compounds", *Ullmann's Encyclopedia of Industrial Chemistry*, Weinheim: Wiley-VCH, doi:10.1002/14356007.a15\_595.pub2

Magnesium bicarbonate	
<b>IUPAC name</b> <span>[hide]</span>	Magnesium hydrogen carbonate
<b>Other names</b> <span>[hide]</span>	Magnesium bicarbonate
Identifiers	
PubChem	102204
ChemSpider	92335 <span>✓</span>
Jmol-3D images	Image 1 <span><span><span></span></span></span>
<b>SMILES</b> <span>[show]</span>	
Properties	
Molecular formula	Mg(HCO <sub>3</sub> ) <sub>2</sub>
Molar mass	146.33868 g/mol
Related compounds	
Other cations	Calcium bicarbonate
Except where noted otherwise, data are given for materials in their <b>standard state</b> (at 25 <span> </span> °C (77 <span> </span> °F), 100 <span> </span> kPa)	
<span>✓</span> <span>(verify)</span> <span>(what is: <span>✓</span>/<span>✗</span>?)</span>	
<b>Infobox references</b>	

v • t • e •

**Magnesium compounds**

[show]

Categories: Magnesium compounds | Bicarbonates