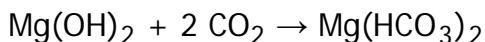


Magnesium bicarbonate

From Wikipedia, the free encyclopedia

Magnesium bicarbonate or **magnesium hydrogen carbonate**, $\text{Mg}(\text{HCO}_3)_2$, is the **bicarbonate** salt of **magnesium**. It can be formed through the reaction of dilute solutions of **carbonic acid** (such as **seltzer water**) and **magnesium hydroxide** (**milk of magnesia**).

Magnesium bicarbonate exists only in aqueous solution. To produce it, a suspension of magnesium hydroxide is treated with pressurized **carbon dioxide**, producing a solution of magnesium bicarbonate: [1]



Drying the resulting solution causes the magnesium bicarbonate to **decompose**, yielding **magnesium carbonate**, carbon dioxide, and water:



References [edit]

1. ^ Margarete Seeger; Walter Otto; Wilhelm Flick; Friedrich Bickelhaupt; Otto S. Akkerman (2005), "Magnesium Compounds", *Ullmann's Encyclopedia of Industrial Chemistry*, Weinheim: Wiley-VCH, doi:10.1002/14356007.a15_595.pub2

Magnesium bicarbonate	
IUPAC name	[hide]
Magnesium hydrogen carbonate	
Other names	[hide]
Magnesium bicarbonate	
Identifiers	
PubChem	102204
ChemSpider	92335 ✓
Jmol-3D images	Image 1
SMILES	[show]
Properties	
Molecular formula	$\text{Mg}(\text{HCO}_3)_2$
Molar mass	146.33868 g/mol
Related compounds	
Other cations	Calcium bicarbonate
Except where noted otherwise, data are given for materials in their standard state (at 25 °C (77 °F), 100 kPa)	
✓ (verify) (what is: ✓ / ✗ ?) Infobox references	

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Magnesium compounds

[show]

Categories: Magnesium compounds | Bicarbonates