

Input interpretation:

1 mM of sodium bicarbonate pH

Result:

6.62

Acid-base information:

K_a	4.69×10^{-11}
pK_a	10.3
pH	6.62
$[H_3O^+]$	$2.38 \times 10^{-7} \text{ mol/L}$ (moles per liter)
pOH	7.38
$[OH^-]$	$4.19 \times 10^{-8} \text{ mol/L}$ (moles per liter)
% ionization	0.0238%

[Source information >](#)