



pH testing

Important Warning

This information is only for patients of Michael Cheikin MD. Risk of serious injury if utilized without proper ongoing medical supervision.

pH is a measure of acid-base, where neutral is 7 (water), 1 is the strongest acid (such as hydrochloric acid), and 14 is the strongest base (such as lye). Blood pH is approximately 7.4. In order to maintain a very tight blood pH, the body excretes acid (or base) in the urine, stool and by other means (such as changing carbon dioxide content via breathing rate). The pH of the stomach during digestion is between 1 and 2 (very acid), while the contents of the small intestine is above 7, ie alkaline. Therefore, stating that we are “acidic” or trying to get “alkaline” are over-simplifications.

Testing urine and salivary pH are indirect measures of bodily processes. While not diagnostic in themselves, observing patterns and trends can be helpful in conjunction with changes in diet, supplements, exercise, sleep, breathing patterns, lab testing etc.

How to use pH Strips: Put a single drop or two on the strip and read immediately.

When to test pH: Best times are first thing in the morning upon awakening, and mid-afternoon before dinner.

How to test **urine** pH: Put a single drop of mid-stream (not initial) urine on the strip. The strip should be put in trash (i.e. don't flush).

How to test **salivary** pH. Upon awakening, you can take a small sip of water if your mouth is dry. Collect saliva in your mouth and swallow x 2, then the third time put a drop of saliva on the pH strip. If later in the day, wait for at least 30 minutes after eating. Rinse mouth with water. Collect saliva in your mouth and swallow x 2, then the third time put a drop of saliva on the pH strip.

Name:

Month:

Day/Date	awake			pm			night			Note (Back)	
	time	urine pH	salivary pH	agent/dose	time	urine pH	agent/dose	time	urine pH		agent/dose
1											
2											
3											
4											
5											
6											
7											
8											
9											
10											
11											
12											
13											
14											

pH scale:

water 7 (as consumed pH is 6.5 because carbon dioxide dissolves from the air into water to make carbonic acid)

acidic < 7.0, lemon juice 2.0

basic or alkaline > 7.0, ammonia

Interesting pH's:

blood 7.4

ocean water 8.1 (200 years ago 8.2). Caused by minerals especially calcium carbonate (CaCO3). Sea water is a buffer which means it contains molecules especially carbonate (CO3) and bicarbonate (HCO3).

oil and alcohol do not have a pH since pH is predominantly a feature of water.

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