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Solubility table

The table below provides information on the variation of solubility of different substances (mostly inorganic compounds) in water with temperature, at one atmosphere pressure. Units of solubility are given in grams per 100 millilitres of water (g/100 mL), unless shown otherwise. The substances are listed in alphabetical order.

Contents

A

Substance	Formula	0 °C	10 °C	15 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Actinium(III) hydroxide	$Ac(OH)_3$				0.0021								
Aluminium chloride	$AlCl_3$	43.9	44.9		45.8	46.6	47.3		48.1		48.6		49.0
Aluminium fluoride	AlF_3	0.57	0.56		0.67	0.78	0.91		1.1		1.32		1.72
Aluminium hydroxide	$Al(OH)_3$				2.262×10^{-8}								
Aluminium nitrate	$Al(NO_3)_3$	60	66.7		73.9	81.8	88.7	96.0	106	120	132	153	160
Aluminium perchlorate	$Al(ClO_4)_3$	122	128		133								
Aluminium sulfate	$Al_2(SO_4)_3$	31.2	33.5		36.4	40.4	45.8	52.2	59.2	66.2	73	80.8	89.0
Ammonia (ml/ml)	NH_3	1176	900		702	565	428	333	252	188	138	100	88
Ammonium acetate	$NH_4C_2H_3O_2$	102			143		204		311		533		
Ammonium azide	NH_4N_3	16			25.3		37.1						
Ammonium benzoate	$NH_4C_7H_5O_2$		19.6		21.3								83
Ammonium bicarbonate	NH_4HCO_3	11.9	16.1		21.7	28.4	36.6		59.2		109	dec	
Ammonium bromide	NH_4Br	60.6	68.1		76.4	83.2	91.2	99.2	108	117	125	135	145
Ammonium carbonate	$(NH_4)_2CO_3{\cdot}H_2O$	55.8			100				dec				
Ammonium chlorate	NH_4ClO_3				28.7								
Ammonium chloride	NH_4Cl	29.4	33.2		37.2	41.4	45.8	50.4	55.3	60.2	65.6	71.2	77.3
Ammonium hexachloroplatinate	$(NH_4)_2PtCl_6$	0.289	0.374		0.499	0.637	0.815		1.44		2.16	2.61	3.36
Ammonium chromate	$(NH_4)_2CrO_4$	25	29.2		34	39.3	45.3	51.9	59.0	71.2	76.1		
Ammonium dichromate	$(NH_4)_2Cr_2O_7$	18.2	25.5		35.6	46.5	58.5	71.4	86.0		115		156
Ammonium dihydrogen arsenate	$NH_4H_2AsO_4$	33.7			48.7		63.8		83		107	122	
Ammonium dihydrogen phosphate	$NH_4H_2PO_4$	22.7	29.5		37.4	46.4	56.7	69.0	82.5	98.6	118.3	142.8	173.2
Ammonium fluoride	NH_4F	71.5	76.1		80.8	86.2	91.57	97.2	103.7	110.5	117.9		
Ammonium fluorosilicate	$(NH_4)_2SiF_6$	12.28	16.41		18.6	25.0	31.6	35.4	40.4	44.9	48.1(75°C)		61.0
Ammonium formate	NH_4HCO_2	102			143		204		311		533		
Ammonium hydrogen phosphate	$(NH_4)_2HPO_4$	42.9	62.9		68.9	75.1	81.8	89.2	97.2	106	110	112	121
Ammonium hydrogen sulfate	NH_4HSO_4				100								
Ammonium hydrogen tartrate	$NH_4HC_4H_4O_6$		1.88		2.7								
Ammonium iodate	NH_4IO_3			2.6									14.5

Ammonium iodide	NH_4I	155	163		172	182	191	200	209	219	229		250
Ammonium nitrate	NH_4NO_3	118	150		192	242	297	344	421	499	580	740	871
Ammonium orthoperiodate	$(NH_4)_5IO_6$				2.7								
Ammonium oxalate	$(NH_4)_2C_2O_4$	2.2	3.21		4.45	6.09	8.18	10.3	14.0		22.4	27.9	34.7
Ammonium perchlorate	NH_4ClO_4	11.56	16.4		20.85		30.58		39.05		48.19		57.01
Ammonium permanganate	NH_4MnO_4			8.0					dec				
Ammonium perrhenate	NH_4ReO_4	2.8			6.2		12.0		20.7		32.3	39.1	
Ammonium phosphate	$(NH_4)_3PO_4$	9.40			20.3			37.7					
Ammonium selenate	$(NH_4)_2SeO_4$	96	105		115	126	143		192				
Ammonium sulfate	$(NH_4)_2SO_4$	70.6	73		75.4	78.1	81.2	84.3	87.4		94.1		103
Ammonium aluminium sulfate	$NH_4Al(SO_4)_2{\cdot}12H_2O$	2.4	5.0		7.4	10.5	14.6	19.6	26.7	37.7	53.9	98.2	121
Ammonium sulfite	$(NH_4)_2SO_3$	47.9	54		60.8	68.8	78.4		104		144	150	153
Ammonium tartrate	$(NH_4)_2C_4H_4O_6$	45	55		63	70.5	76.5		86.9				
Ammonium thiocyanate	NH_4SCN	120	144		170	208	234	235	346				
Ammonium thiosulfate	$(NH_4)_2S_2O_3$				173		205				269		
Ammonium vanadate	NH_4VO_3				0.48	0.84	1.32	1.78	2.42	3.05			7.0
Aniline	C_6H_7N				3.6								
Antimony trifluoride	SbF_3	385			444	562	dec						
Antimony sulfide	Sb_2S_3				1.8×10^{-4}								
Antimony trichloride	$SbCl_3$	602			910	1090	1370	1917	4531	dec			
Argon (Unit:ml/ml)	Ar	0.056	0.0405		0.0336	0.0288	0.0252	0.0223					
Arsenic pentasulfide	As_2S_5	0.0014											
Arsenic pentoxide	As_2O_5	59.5	62.1		65.8	70.6	71.2		73.0		75.1		76.7
Arsenious sulfide	As_2S_3				0.0004								
Arsenic trioxide	As_2O_3	1.21	1.58		1.80		2.93	3.43	4.44	5.37	5.89	6.55	9
Arsine (Unit:ml/ml)	AsH_3				0.2								

B

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Barium acetate	$Ba(C_2H_3O_2)_2$	58.8	62	72	75	78.5	77	75	74	74		
Barium arsenate	$Ba_3(AsO_4)_2$			2.586×10^{-9}								
Barium azide	$Ba(N_3)_2$	12.5	16.1	17.4					24.75			
Barium bromate	$Ba(BrO_3)_2{\cdot}H_2O$	0.29	0.44	0.65	0.95	1.31	1.75	2.27	3.01	3.65	4.45	5.71
Barium bromide	$BaBr_2$	98	101	104	109	114		123		135		149
Barium carbonate	$BaCO_3$			2.4×10^{-3}								
Barium chlorate	$Ba(ClO_3)_2$	20.3	26.9	33.9	41.6	49.7		66.7		84.8		105
Barium chloride	$BaCl_2$	31.2	33.5	35.8	38.1	40.8		46.2		52.5	55.8	59.4
Barium chlorite	$Ba(ClO_2)_2$	43.9	44.6	45.4		47.9		53.8		66.6		80.8
Barium chromate	$BaCrO_4$			2.775×10^{-4}								
Barium cyanide	$Ba(CN)_2$			80								

Barium ferrocyanide	$Ba_2Fe(CN)_6$			0.009732								
Barium fluoride	BaF_2		0.159	0.16	0.161							
Barium fluorosilicate	$BaSiF_6$			0.028								
Barium formate	$Ba(HCO_2)_2$	26.2	28	31.9	34		38.6		44.2	47.6	51.3	
Barium hydrogen phosphate	$BaHPO_4$			0.013								
Barium hydrogen phosphite	$BaHPO_3$			0.687								
Barium hydroxide	$Ba(OH)_2{\cdot}8H_2O$	1.67	2.48	3.89	5.59	8.22	11.7	20.9		101		
Barium iodate	$Ba(IO_3)_2$			0.035	0.046	0.057						0.2
Barium iodide	BaI_2	182	201	223	250			264			291	301
Barium molybdate	$BaMoO_4$			0.006								
Barium nitrate	$Ba(NO_3)_2$	4.95	6.77	9.02	11.5	14.1		20.4		27.2		34.4
Barium nitrite	$Ba(NO_2)_2$	50.3	60	72.8		102		151		222	261	325
Barium oxalate	$BaC_2O_4{\cdot}2H_2O$			0.003								
Barium oxide	BaO			3.48							90.8	
Barium perchlorate	$Ba(ClO_4)_2$	239		336		416		495		575		653
Barium permanganate	$Ba(MnO_4)_2$		62.5									
Barium manganate	$BaMnO_4$			0.0036								
Barium pyrophosphate	$Ba_2P_2O_7$			0.009								
Barium selenate	$BaSeO_4$			0.005								
Barium sulfate	$BaSO_4$			2.448×10^{-4}	2.85×10^{-4}							
Beryllium carbonate	$BeCO_3$			0.218								
Beryllium chloride	$BeCl_2$		42	42								
Beryllium molybdate	$BeMoO_4$			3.02								
Beryllium nitrate	$Be(NO_3)_2$	97	102	108	113	125		178				
Beryllium oxalate	$BeC_2O_4{\cdot}3H_2O$			63.5								
Beryllium perchlorate	$Be(ClO_4)_2$			147								
Beryllium selenate	$BeSeO_4{\cdot}4H_2O$			49								
Beryllium sulfate	$BeSO_4{\cdot}4H_2O$	37	37.6	39.1	41.4	45.8		53.1		67.2		82.8
Bismuth arsenate	$BiAsO_4$			7.298×10^{-4}								
Bismuth hydroxide	$Bi(OH)_3$			2.868×10^{-7}								
Bismuth iodide	BiI_3			7.761×10^{-4}								
Bismuth phosphate	$BiPO_4$			1.096×10^{-10}								
Bismuth sulfide	Bi_2S_3			1.561×10^{-20}								
Boric acid	H_3BO_3	2.52	3.49	4.72	6.23	8.08	10.27	12.97	15.75	19.10	23.27	27.53
Boron trioxide	B_2O_3			2.2								
Bromine monochloride	$BrCl$			1.5								

C

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Cadmium arsenate	$Cd_3(AsO_4)_2$			7.091×10^{-6}								
Cadmium benzoate	$Cd(C_7H_5O_2)_2$			2.81								
Cadmium bromate	$Cd(BrO_3)_2$			125								
Cadmium bromide	$CdBr_2$	56.3	75.4	98.8	129	152		153		156		160
Cadmium carbonate	$CdCO_3$			3.932×10^{-5}								

Cadmium chlorate	$Cd(ClO_3)_2$	299	308	322	348	376		455				
Cadmium chloride	$CdCl_2$	100	135	135	135	135		136		140		147
Cadmium cyanide	$Cd(CN)_2$			0.022								
Cadmium ferrocyanide	$Cd_2Fe(CN)_6$			8.736×10^{-5}								
Cadmium fluoride	CdF_2			4								
Cadmium formate	$Cd(HCO_2)_2$	8.3	11.1	14.4	18.6	25.3		59.5		80.5	85.2	94.6
Cadmium hydroxide	$Cd(OH)_2$			2.697×10^{-4}								
Cadmium iodate	$Cd(IO_3)_2$			0.097								
Cadmium iodide	CdI_2	78.7		84.7	87.9	92.1		100		111		125
Cadmium nitrate	$Cd(NO_3)_2$	122		136	150	194		310		713		
Cadmium oxalate	$CdC_2O_4{\cdot}3H_2O$			0.006046								
Cadmium perchlorate	$Cd(ClO_4)_2$		180	188	195	203		221		243		272
Cadmium phosphate	$Cd_3(PO_4)_2$			6.235×10^{-6}								
Cadmium selenate	$CdSeO_4$	72.5	68.4	64	58.9	55		44.2		32.5	27.2	22
Cadmium sulfate	$CdSO_4$	75.4	76	76.6		78.5		81.8		66.7	63.1	60.8
Cadmium sulfide	CdS			1.292×10^{-12}								
Cadmium tungstate	$CdWO_4$			0.04642								
Caesium acetate	$CsC_2H_3O_2$			1010							1345.5	
Caesium azide	CsN_3			307								
Caesium bromate	$CsBrO_3$	2.10		3.66	4.53	5.3						
Caesium bromide	$CsBr$			108								
Caesium chlorate	$CsClO_3$		3.8	6.2	9.5	13.8		26.2		45	58	79
Caesium chloride	$CsCl$	146	175	187	197	208		230		250	260	271
Caesium chromate	Cs_2CrO_4		71.4									
Caesium fluoride	CsF			322								
Caesium fluoroborate	$CsBF_4$			0.818								
Caesium formate	$CsHCO_2$	335	381	450	694							
Caesium iodate	$CsIO_3$			2.6								
Caesium iodide	CsI	44.1	58.5	76.5	96	124		150		190	205	
Caesium nitrate	$CsNO_3$	9.33	14.9	23	33.9	47.2		83.8		134	163	197
Caesium oxalate	$Cs_2C_2O_4$			313								
Caesium perchlorate	$CsClO_4$	0.8	1	1.6	2.6	4		7.3		14.4	20.5	30
Caesium permanganate	$CsMnO_4$			0.228								
Caesium phosphate	Cs_3PO_4			340								
Caesium selenate	Cs_2SeO_4		244									
Caesium sulfate	Cs_2SO_4	167	173	179	184	190		200		210	215	200
Calcium acetate	$Ca(C_2H_3O_2)_2$	37.4	36	34.7	33.8	33.2		32.7		33.5	31.1	29.7
Calcium arsenate	$Ca_3(AsO_4)_2$			0.003629								
Calcium azide	$Ca(N_3)_2$			45								
Calcium benzoate	$Ca(C_7H_5O_2)_2{\cdot}3H_2O$	2.32	2.45	2.72	3.02	3.42		4.71		6.87	8.55	8.7
Calcium bicarbonate	$Ca(HCO_3)_2$	16.1		16.6		17.1		17.5		17.9		18.4
Calcium bromate	$Ca(BrO_3)_2$			230								
Calcium bromide	$CaBr_2$	125	132	143		213		278		295		312
Calcium carbonate (Aragonite)	$CaCO_3$-Aragonite			7.753×10^{-4}								

Calcium carbonate (Calcite)	$CaCO_3$-Calcite			6.17×10^{-4}								
Calcium chlorate	$Ca(ClO_3)_2$			209								
Calcium chloride	$CaCl_2$	59.5	64.7	74.5	100	128		137		147	154	159
Calcium chromate	$CaCrO_4$	4.5		2.25	1.83	1.49		0.83				
Calcium citrate	$Ca_3(C_6H_5O_7)_2$			0.095 **(25 °C)**								
Monocalcium phosphate	$Ca(H_2PO_4)_2$			1.8								
Calcium fluoride	CaF_2	0.008575										
Calcium fluorosilicate	$CaSiF_6$			0.518								
Calcium formate	$Ca(HCO_2)_2$	16.1		16.6		17.1		17.5		17.9		18.4
Dicalcium phosphate	$CaHPO_4$			0.004303								
Calcium hydroxide	$Ca(OH)_2$	0.189	0.182	0.173	0.16	0.141		0.121		0.086	0.076	0.068
Calcium iodate	$Ca(IO_3)_2$	0.09		0.24	0.38	0.52		0.65		0.66	0.67	0.67
Calcium iodide	CaI_2	64.6		66	67.6	70.8		74		78		81
Calcium molybdate	$CaMoO_4$			0.004099								
Calcium nitrate	$Ca(NO_3)_2$			121.2								
Calcium nitrate	$Ca(NO_3)_2{\cdot}4H_2O$	102	115	129	152	191				358		363
Calcium nitrite	$Ca(NO_2)_2{\cdot}4H_2O$	63.9		84.5	104			134		151	166	178
Calcium oxalate	CaC_2O_4			6.7×10^{-4}							0.0014	
Calcium oxide	CaO											5.7
Calcium perchlorate	$Ca(ClO_4)_2$			188								
Calcium permanganate	$Ca(MnO_4)_2$			338								
Calcium phosphate	$Ca_3(PO_4)_2$			0.002								
Calcium selenate	$CaSeO_4{\cdot}2H_2O$	9.73	9.77	9.22	8.79	7.14						
Calcium sulfate	$CaSO_4{\cdot}2H_2O$	0.223	0.244	0.255	0.264	0.265		0.244		0.234		0.205
Calcium tungstate	$CaWO_4$			0.002387								
Carbon dioxide	CO_2			0.1782								
Carbon monoxide	CO			0.0026								
Cerium(III) acetate	$Ce(C_2H_3O_2)_3$			0.35[1]								
Cerium(III) chloride	$CeCl_3$			100								
Cerium(III) hydroxide	$Ce(OH)_3$			9.43×10^{-5}								
Cerium(III) iodate	$Ce(IO_3)_3$			0.123								
Cerium(III) nitrate	$Ce(NO_3)_3$			234								
Cerium(III) phosphate	$CePO_4$			7.434×10^{-11}								
Cerium(III) selenate	$Ce_2(SeO_4)_3$	39.5	37.2	35.2	33.2	32.6		13.7		4.6		
Cerium(III) sulfate	$Ce_2(SO_4)_3{\cdot}2H_2O$	21.4		9.84	7.24	5.63		3.87				
Cerium(IV) hydroxide	$Ce(OH)_4$			1.981×10^{-5}								
Chromium(III) nitrate	$Cr(NO_3)_3$	108	124	130	152							
Chromium(III) perchlorate	$Cr(ClO_4)_3$	104	123	130								
Chromium(III) sulfate	$Cr_2(SO_4)_3{\cdot}18H_2O$			220								
Chromium(VI) oxide	CrO_3	61.7		63							67	
Cobalt(II) bromate	$Co(BrO_3)_2{\cdot}6H_2O$			45.5								
Cobalt(II) bromide	$CoBr_2$	91.9		112	128	163		227		241		257
Cobalt(II) chlorate	$Co(ClO_3)_2$	135	162	180	195	214		316				
Cobalt(II) chloride	$CoCl_2$	43.5	47.7	52.9	59.7	69.5		93.8		97.6	101	106

Cobalt(II) fluoride	CoF_2			1.36								
Cobalt(II) fluorosilicate	$CoSiF_6{\cdot}6H_2O$			118								
Cobalt(II) iodate	$Co(IO_3)_2{\cdot}2H_2O$			1.02	0.9	0.88		0.82		0.73		0.7
Cobalt(II) iodide	CoI_2			203								
Cobalt(II) nitrate	$Co(NO_3)_2$	84	89.6	97.4	111	125		174		204	300	
Cobalt(II) nitrite	$Co(NO_2)_2$	0.076	0.24	0.4	0.61	0.85						
Cobalt oxalate	$CoC_2O_4{\cdot}2H_2O$			2.6972×10^{-9}								
Cobalt(II) perchlorate	$Co(ClO_4)_2$			104								
Cobalt(II) sulfate	$CoSO_4$	25.5	30.5	36.1	42	48.8		55		53.8	45.3	38.9
Copper(I) chloride	CuCl			0.0099								
Copper(I) cyanide	CuCN			1.602×10^{-9}								
Copper(I) hydroxide	CuOH			8.055×10^{-7}								
Copper(I) iodide	CuI			0.0042								
Copper(I) sulfide	Cu_2S			1.361×10^{-15}								
Copper(I) thiocyanate	CuSCN			8.427×10^{-7}								
Copper(II) bromide	$CuBr_2$	107	116	126	128	131						
Copper(II) carbonate	$CuCO_3$			1.462×10^{-4}								
Copper(II) chlorate	$Cu(ClO_3)_2$			242								
Copper(II) chloride	$CuCl_2$	68.6	70.9	73	77.3	87.6		96.5		104	108	120
Copper(II) chromate	$CuCrO_4$			0.03407								
Copper(II) fluoride	CuF_2			0.075								
Copper(II) fluorosilicate	$CuSiF_6$	73.5	76.5	81.6	84.1	91.2				93.2		
Copper(II) formate	$Cu(HCO_2)_2$			12.5								
Copper(II) hydroxide	$Cu(OH)_2$			1.722×10^{-6}								
Copper(II) iodate	$Cu(IO_3)_2{\cdot}2H_2O$			0.109								
Copper(II) nitrate	$Cu(NO_3)_2$	83.5	100	125	156	163		182		208	222	247
Copper oxalate	$CuC_2O_4{\cdot}2H_2O$			2.1627×10^{-10}								
Copper(II) perchlorate	$Cu(ClO_4)_2$				146							
Copper(II) selenate	$CuSeO_4$	12	14.5	17.5	21	25.2		36.5		53.7		
Copper(II) selenite	$CuSeO_3$			0.002761								
Copper(II) sulfate	$CuSO_4{\cdot}5H_2O$	23.1	27.5	32	37.8	44.6		61.8		83.8		114
Copper(II) sulfide	CuS			2.41×10^{-17}								

D and E

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Dysprosium(III) chromate	$Dy_2(CrO_4)_3{\cdot}10H_2O$			0.663								
Dysprosium(III) sulfate	$Dy_2(SO_4)_3{\cdot}8H_2O$			4.83								
Erbium(III) hydroxide	$Er(OH)_3$			1.363×10^{-5}								
Erbium(III) sulfate	$Er_2(SO_4)_3$			13.79								
Erbium(III) sulfate	$Er_2(SO_4)_3{\cdot}8H_2O$			16.00		6.53						
Europium(III) hydroxide	$Eu(OH)_3$			1.538×10^{-5}								
Europium(III) sulfate	$Eu_2(SO_4)_3{\cdot}8H_2O$			2.56								

F and G

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Ferrous ammonium sulfate	$(NH_4)_2Fe(SO_4)_2·6H_2O$			26.9						73		
Fructose	$C_6H_{12}O_6$			375.0		538.0						
Gadolinium(III) acetate	$Gd(C_2H_3O_2)_3·4H_2O$			11.6								
Gadolinium(III) bicarbonate	$Gd(HCO_3)_3$			5.61								
Gadolinium(III) bromate	$Gd(BrO_3)_3·9H_2O$	50.2	70.1	95.6	126	166						
Gadolinium(III) hydroxide	$Gd(OH)_3$			$1.882×10^{-5}$								
Gadolinium(III) sulfate	$Gd_2(SO_4)_3$	3.98	3.3	2.6	2.32							
D-Galactose	$C_6H_{12}O_6$			10.3								68.3
Gallium chloride	$GaCl_3$			180								
Gallium hydroxide	$Ga(OH)_3$			$8.616×10^{-9}$								
Gallium oxalate	$Ga_2(C_2O_4)_3·4H_2O$			0.4								
Gallium selenate	$Ga_2(SeO_4)_3·16H_2O$			18.1								
D-Glucose	$C_6H_{12}O_6$			90								
Gold(III) chloride	$AuCl_3$			68								
Gold(V) oxalate	$Au_2(C_2O_4)_5$			0.258								

H

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Hafnium(III) hydroxide	$Hf(OH)_3$			$4.503×10^{-4}$								
Hafnium(IV) hydroxide	$Hf(OH)_4$			$4.503×10^{-6}$								
Helium	He			0.6								
Holmium(III) hydroxide	$Ho(OH)_3$			$2.519×10^{-5}$								
Holmium(III) sulfate	$Ho_2(SO_4)_3·8H_2O$			8.18	6.1	4.52						
Hydrogen chloride	HCl	81	75	70	65.5	61	57.5	53	50	47	43	40
Hydrogen sulfide	H_2S			0.33								

I

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Indium(III) bromide	$InBr_3$			571								
Indium(III) chloride	$InCl_3$		210	212								
Indium(III) fluoride	InF_3			11.2								
Indium(III) hydroxide	$In(OH)_3$			$3.645×10^{-8}$								
Indium(III) iodate	$In(IO_3)_3$			0.067								
Indium(III) sulfide	In_2S_3			$2.867×10^{-14}$								
Iron(II) bromide	$FeBr_2$	101	109	117	124	133		144		168	176	184
Iron(II) carbonate	$FeCO_3$			$6.554×10^{-5}$								
Iron(II) chloride	$FeCl_2$	49.7	59	62.5	66.7	70		78.3		88.7	92.3	94.9
Iron(II) fluorosilicate	$FeSiF_6·6H_2O$	72.1	74.4		77			84		88		100
Iron(II) hydroxide	$Fe(OH)_2$			$5.255×10^{-5}$								
Iron(II) nitrate	$Fe(NO_3)_2·6H_2O$	113	134									
Iron(II) oxalate	$FeC_2O_4·2H_2O$			0.008								

Iron(II) perchlorate	$Fe(ClO_4)_2·6H_2O$			299								
Iron(II) sulfate	$FeSO_4$			28.8		40	48	60	73.3		101	79.9
Iron(III) arsenate	$FeAsO_4$			$1.47×10^{-9}$								
Iron(III) chloride	$FeCl_3·6H_2O$	74.4		91.8	107							
Iron(III) fluoride	FeF_3			0.091								
Iron(III) hydroxide	$Fe(OH)_3$			$2.097×10^{-9}$								
Iron(III) iodate	$Fe(IO_3)_3$			0.36								
Iron(III) nitrate	$Fe(NO_3)_3·9H_2O$	112		138		175						
Iron(III) perchlorate	$Fe(ClO_4)_3$	289		368	422	478		772				
Iron(III) sulfate	$Fe_2(SO_4)_3·H_2O$			25.6								

L

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Lactose	$C_{12}H_{22}O_{11}$			8								
Lanthanum(III) acetate	$La(C_2H_3O_2)_3·H_2O$			16.9								
Lanthanum(III) bromate	$La(BrO_3)_3$	98	120	149	200							
Lanthanum(III) carbonate	$La_2(CO_3)_3$			0.0000125								
Lanthanum(III) iodate	$La(IO_3)_3$			0.04575								
Lanthanum(III) molybdate	$La_2(MoO_4)_3$			0.002473								
Lanthanum(III) nitrate	$La(NO_3)_3$	100		136		168		247				
Lanthanum(III) selenate	$La_2(SeO_4)_3$	50.5	45	45	45	45		18.5		5.4	2.2	
Lanthanum(III) sulfate	$La_2(SO_4)_3$	3	2.72	2.33	1.9	1.67		1.26		0.91	0.79	0.69
Lanthanum(III) tungstate	$La_2(WO_4)_3·3H_2O$			6.06								
Lead(II) acetate	$Pb(C_2H_3O_2)_2$	19.8	29.5	44.3	69.8	116						
Lead(II) azide	$Pb(N_3)_2$			0.0249								
Lead(II) bromate	$Pb(BrO_3)_2$			7.92								
Lead(II) bromide	$PbBr_2$	0.45	0.63	0.973	1.12	1.5		2.29		3.32	3.86	4.55
Lead(II) carbonate	$PbCO_3$			$7.269×10^{-5}$								
Lead(II) chlorate	$Pb(ClO_3)_2$			144								
Lead(II) chloride	$PbCl_2$	0.67	0.82	1.08	1.2	1.42		1.94		2.54	2.88	3.2
Lead(II) chromate	$PbCrO_4$			$1.71×10^{-5}$								
Lead(II) ferrocyanide	$PbFe(CN)_6$			$5.991×10^{-4}$								
Lead(II) fluoride	PbF_2			0.0671								
Lead(II) fluorosilicate	$PbSiF_6$	190		222				403		428		463
Lead(II) hydrogen phosphate	$PbHPO_4$			$3.457×10^{-4}$								
Lead(II) hydrogen phosphite	$PbHPO_3$			0.02187								
Lead(II) hydroxide	$Pb(OH)_2$			$1.615×10^{-4}$								
Lead(II) iodate	$Pb(IO_3)_2$			0.0024								
Lead(II) iodide	PbI_2	0.044	0.056	0.076	0.09	0.124		0.193		0.294		0.42
Lead(II) molybdate	$PbMoO_4$			$1.161×10^{-5}$								
Lead(II) nitrate	$Pb(NO_3)_2$	37.5	46.2	54.3	63.4	72.1		91.6		111		133
Lead(II) oxalate	PbC_2O_4			$6.495×10^{-4}$								

Lead(II) perchlorate	$Pb(ClO_4)_2·3H_2O$			440								
Lead(II) selenate	$PbSeO_4$			0.0131								
Lead(II) sulfate	$PbSO_4$			0.00443								
Lead(II) sulfide	PbS			$6.767×10^{-13}$								
Lead(II) tartrate	$PbC_4H_4O_6$			0.0025								
Lead(II) thiocyanate	$Pb(SCN)_2$			0.553								
Lead(II) thiosulfate	PbS_2O_3			0.0202								
Lead(II) tungstate	$PbWO_4$			0.02838								
Lead(IV) hydroxide	$Pb(OH)_4$			$7.229×10^{-11}$								
Lithium acetate	$LiC_2H_3O_2$	31.2	35.1	40.8	50.6	68.6						
Lithium azide	LiN_3	61.3	64.2	67.2	71.2	75.4		86.6				100
Lithium benzoate	$LiC_7H_5O_2$	38.9	41.6	44.7	53.8							
Lithium bicarbonate	$LiHCO_3$			5.74								
Lithium bromate	$LiBrO_3$	154	166	179	198	221		269		308	329	355
Lithium bromide	$LiBr$	143	147	160	183	211		223		245		266
Lithium carbonate	Li_2CO_3	1.54	1.43	1.33	1.26	1.17		1.01		0.85		0.72
Lithium chlorate	$LiClO_3$	241	283	372	488	604		777				
Lithium chloride	$LiCl$	69.2	74.5	83.5	86.2	89.8		98.4		112	121	128
Lithium chromate	$Li_2CrO_4·2H_2O$			142								
Lithium dichromate	$Li_2Cr_2O_7·2H_2O$				151							
Lithium dihydrogen phosphate	LiH_2PO_4	126										
Lithium fluoride	LiF			0.127	0.135							
Lithium fluorosilicate	$Li_2SiF_6·2H_2O$			73								
Lithium formate	$LiHCO_2$	32.3	35.7	39.3	44.1	49.5		64.7		92.7	116	138
Lithium hydrogen phosphite	Li_2HPO_3	4.43			9.97	7.61		7.11				6.03
Lithium hydroxide	$LiOH$	12.7	12.7	12.8	12.9	13.0	13.3	13.8		15.3		17.5
Lithium iodide	LiI	151	157	165	171	179		202		435	440	481
Lithium molybdate	Li_2MoO_4	82.6		79.5	79.5	78						73.9
Lithium nitrate	$LiNO_3$	53.4	60.8	70.1	138	152		175				
Lithium nitrite	$LiNO_2$	70.9	82.5	96.8	114	133		177		233	272	324
Lithium oxalate	$Li_2C_2O_4$			8								
Lithium perchlorate	$LiClO_4$	42.7	49	56.1	63.6	72.3		92.3		128	151	
Lithium permanganate	$LiMnO_4$			71.4								
Lithium phosphate	Li_3PO_4			0.039								
Lithium selenide	Li_2Se			57.7								
Lithium selenite	Li_2SeO_3	25	23.3	21.5	19.6	17.9		14.7		11.9	11.1	9.9
Lithium sulfate	Li_2SO_4	36.1	35.5	34.8	34.2	33.7		32.6		31.4	30.9	
Lithium tartrate	$Li_2C_4H_4O_6$	42	31.8	27.1	26.6	27.2		29.5				
Lithium thiocyanate	$LiSCN$			114	131	153						
Lithium vanadate	$LiVO_3$	2.5		4.82	6.28	4.38		2.67				
Lutetium(III) hydroxide	$Lu(OH)_3$			$1.164×10^{-5}$								
Lutetium(III) sulfate	$Lu_2(SO_4)_3·8H_2O$			57.9								

M

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C

Magnesium acetate	$Mg(C_2H_3O_2)_2$	56.7	59.7	53.4	68.6	75.7		118				
Magnesium benzoate	$Mg(C_7H_5O_2)_2{\cdot}H_2O$					5						
Magnesium bromate	$Mg(BrO_3)_2{\cdot}6H_2O$					58						
Magnesium bromide	$MgBr_2$	98	99	101	104	106		112				125
Magnesium carbonate	$MgCO_3$			0.039								
Magnesium chlorate	$Mg(ClO_3)_2$	114	123	135	155	178		242			268	
Magnesium chloride	$MgCl_2$	52.9	53.6	54.6	55.8	57.5		61		66.1	69.5	73.3
Magnesium chromate	$MgCrO_4{\cdot}7H_2O$			137								
Magnesium fluoride	MgF_2			0.007325								
Magnesium fluorosilicate	$MgSiF_6$	26.3		30.8		34.9		44.4				
Magnesium formate	$Mg(HCO_2)_2$	14	14.2	14.4	14.9	15.9		17.9		20.5	22.2	22.9
Magnesium hydroxide	$Mg(OH)_2$			9.628×10^{-4}								0.004
Magnesium iodate	$Mg(IO_3)_2$		7.2	8.6	10	11.7		15.2		15.5	15.6	
Magnesium iodide	MgI_2	120		140		173				186		
Magnesium molybdate	$MgMoO_4$			13.7								
Magnesium nitrate	$Mg(NO_3)_2$	62.1	66	69.5	73.6	78.9		78.9		91.6	106	
Magnesium oxalate	MgC_2O_4			0.104								
Magnesium oxide	MgO			0.009								
Magnesium perchlorate	$Mg(ClO_4)_2$			49.6								
Magnesium phosphate	$Mg_3(PO_4)_2$			2.588×10^{-4}								
Magnesium selenate	$MgSeO_4$	20	30.4	38.3	44.3	48.6		55.8				
Magnesium selenite	$MgSeO_3$			0.05454								
Magnesium sulfate	$MgSO_4$	25.5	30.4	35.1	39.7	44.7	50.4	54.8	59.2	54.8	52.9	50.2
Magnesium sulfite	$MgSO_3{\cdot}6H_2O$			0.52**(25 °C)**[2]								
Magnesium thiosulfate	MgS_2O_3			50								
Maltose	$C_{12}H_{22}O_{11}$			108								
D-Mannose	$C_6H_{12}O_6$			248								
Manganese(II) bromide	$MnBr_2$	127	136	147	157	169		197		225	226	228
Manganese(II) carbonate	$MnCO_3$			4.877×10^{-5}								
Manganese(II) chloride	$MnCl_2$	63.4	68.1	73.9	80.8	88.5		109		113	114	115
Manganese(II) ferrocyanide	$Mn_2Fe(CN)_6$			0.001882								
Manganese(II) fluoride	MnF_2			0.96		0.67		0.44				0.48
Manganese(II) fluorosilicate	$MnSiF_6{\cdot}6H_2O$			140								
Manganese(II) hydroxide	$Mn(OH)_2$			3.221×10^{-4}								
Manganese(II) nitrate	$Mn(NO_3)_2$	102	118	139	206							
Manganese(II) oxalate	$MnC_2O_4{\cdot}2H_2O$	0.02	0.024	0.028	0.033							
Manganese(II) sulfate	$MnSO_4$	52.9	59.7	62.9	62.9	60		53.6		45.6	40.9	35.3
Mercury(I) azide	$Hg_2(N_3)_2$			0.02727								
Mercury(I) bromide	Hg_2Br_2			1.352×10^{-6}								
Mercury(I) carbonate	Hg_2CO_3			4.351×10^{-7}								
Mercury(I) chloride	Hg_2Cl_2			3.246×10^{-5}								
Mercury(I) chromate	Hg_2CrO_4			0.002313								
Mercury(I) cyanide	$Hg_2(CN)_2$			2.266×10^{-12}								

Mercury(I) perchlorate	$Hg_2(ClO_4)_2$	282	325	407	455		499		541		580	
Mercury(I) sulfate	Hg_2SO_4			0.04277								
Mercury(II) acetate	$Hg(C_2H_3O_2)_2$			25								
Mercury(II) benzoate	$Hg(C_7H_5O_2)_2·H_2O$			1.1								
Mercury(II) bromate	$Hg(BrO_3)_2·2H_2O$			0.08								
Mercury(II) bromide	$HgBr_2$	0.3	0.4	0.56	0.66	0.91		1.68		2.77		4.9
Mercury(II) chlorate	$Hg(ClO_3)_2$			25								
Mercury(II) chloride	$HgCl_2$	3.63	4.82	6.57	8.34	10.2		16.3		30		61.3
Mercury(II) cyanide	$Hg(CN)_2$			9.3								
Mercury(II) iodate	$Hg(IO_3)_2$			0.002372								
Mercury(II) iodide	HgI_2			0.006								
Mercury(II) oxalate	HgC_2O_4			0.011								
Mercury(II) sulfide	HgS			$2.943×10^{-25}$								
Mercury(II) thiocyanate	$Hg(SCN)_2$			0.063								

N and O

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Neodymium(III) acetate	$Nd(C_2H_3O_2)_3·H_2O$			26.2								
Neodymium(III) bromate	$Nd(BrO_3)_3$	43.9	59.2	75.6	95.2	116						
Neodymium(III) chloride	$NdCl_3$		96.7	98	99.6	102		105				
Neodymium(III) molybdate	$Nd_2(MoO_4)_3$				0.0019							
Neodymium(III) nitrate	$Nd(NO_3)_3$	127	133	142	145	159		211				
Neodymium(III) selenate	$Nd_2(SeO_4)_3$	45.2	44.6	41.8	39.9	39.9		43.9		7	3.3	
Neodymium(III) sulfate	$Nd_2(SO_4)_3$	13	9.7	7.1	5.3	4.1		2.8		2.2	1.2	
Nickel(II) acetate	$Ni(C_2H_3O_2)_2$											
Nickel(II) bromate	$Ni(BrO_3)_2·6H_2O$			28								
Nickel(II) bromide	$NiBr_2$	113	122	131	138	144		153		154		155
Nickel(II) carbonate	$NiCO_3$			$9.643×10^{-4}$								
Nickel(II) chlorate	$Ni(ClO_3)_2$	111	120	133	155	181		221		308		
Nickel(II) chloride	$NiCl_2$	53.4	56.3	66.8	70.6	73.2		81.2		86.6		87.6
Nickel(II) fluoride	NiF_2		2.55	2.56				2.56			2.59	
Nickel(II) formate	$Ni(HCO_2)_2.2·H_2O$		3.15	3.25								
Nickel(II) hydroxide	$Ni(OH)_2$			0.013								
Nickel(II) iodate	$Ni(IO_3)_2$	0.74		0.062	1.43							
Nickel(II) iodide	NiI_2	124	135	148	161	174		184		187	188	
Nickel(II) nitrate	$Ni(NO_3)_2$	79.2		94.2	105	119		158		187	188	
Nickel oxalate	$NiC_2O_4·2H_2O$			0.00118								
Nickel(II) perchlorate	$Ni(ClO_4)_2$	105	107	110	113	117						
Nickel(II) pyrophosphate	$Ni_2P_2O_7$			0.001017								
Nickel(II) sulfate	$NiSO_4·6H_2O$			44.4	46.6	49.2		55.6		64.5	70.1	76.7
Nitric oxide	NO			0.0056								
Nitrous oxide	N_2O			0.112								

Oxygen at a partial pressure of 21 kPa	O_2	0.00146	0.00113	0.00091	0.00076	0.00065						
Oxalic acid	$H_2C_2O_4{\cdot}2H_2O$	4.96	8.51	13.3	19.9	30.1		62.1		118	168	

P

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Palladium(II) hydroxide	$Pd(OH)_2$			4.106×10^{-10}								
Palladium(IV) hydroxide	$Pd(OH)_4$			5.247×10^{-14}								
Phenol	C_6H_5OH			8.3		miscible						
Platinum(II) hydroxide	$Pt(OH)_2$			3.109×10^{-11}								
Platinum(IV) bromide	$PtBr_4$			1.352×10^{-7}								
Plutonium(III) fluoride	PuF_3			3.144×10^{-4}								
Plutonium(IV) fluoride	PuF_4			3.622×10^{-4}								
Plutonium(IV) iodate	$Pu(IO_3)_4$			0.07998								
Polonium(II) sulfide	PoS			2.378×10^{-14}								
Potassium acetate	$KC_2H_3O_2$	216	233	256	283	324		350		381	398	
Potassium arsenate	K_3AsO_4			19								
Potassium azide	KN_3	41.4	46.2	50.8	55.8	61						106
Potassium benzoate	$KC_7H_5O_2$		65.8	70.7	76.7	82.1						
Potassium bromate	$KBrO_3$	3.09	4.72	6.91	9.64	13.1		22.7		34.1		49.9
Potassium bromide	KBr	53.6	59.5	65.3	70.7	75.4		85.5		94.9	99.2	104
Potassium hexabromoplatinate	K_2PtBr_6			1.89								
Potassium carbonate	K_2CO_3	105	109	111	114	117	121.2	127		140	148	156
Potassium chlorate	$KClO_3$	3.3	5.2	7.3	10.1	13.9		23.8		37.5	46	56.3
Potassium chloride	KCl	28	31.2	34.2	37.2	40.1	42.6	45.8		51.3	53.9	56.3
Potassium chromate	K_2CrO_4	56.3	60	63.7	66.7	67.8		70.1			74.5	
Potassium cyanide	KCN			50								
Potassium dichromate	$K_2Cr_2O_7$	4.7	7	12.3	18.1	26.3	34	45.6		73		
Potassium dihydrogen arsenate	KH_2AsO_4			19								
Potassium dihydrogen phosphate	KH_2PO_4	14.8	18.3	22.6	28	35.5	41	50.2		70.4	83.5	
Potassium ferricyanide	$K_3Fe(CN)_6$	30.2	38	46	53	59.3		70				91
Potassium ferrocyanide	$K_4Fe(CN)_6$	14.3	21.1	28.2	35.1	41.4		54.8		66.9	71.5	74.2
Potassium fluoride	KF	44.7	53.5	94.9	108	138		142		150		
Potassium formate	$KHCO_2$	328	313	337	361	398		471		580	658	
Potassium hydrogen carbonate	$KHCO_3$	22.5	27.4	33.7	39.9	47.5		65.6				
Potassium hydrogen phosphate	K_2HPO_4			150								
Potassium hydrogen sulfate	$KHSO_4$	36.2		48.6	54.3	61		76.4		96.1		122
Potassium hydrogen tartrate	$KHC_4H_4O_6$			0.6								6,2
Potassium hydroxide	KOH	95.7	103	112	126	134		154				178
Potassium iodate	KIO_3	4.6	6.27	8.08	10.3	12.6	14	18.3		24.8		32.3
Potassium iodide	KI	128	136	144	153	162		176		192	198	206
Potassium metabisulfite	$K_2S_2O_5$				45							
Potassium nitrate	KNO_3	13.3	20.9	31.6	45.8	63.9	85.5	110.0	138	169	202	246

Potassium nitrite	KNO_2	279	292	306	320	329		348		376	390	410
Potassium oxalate	$K_2C_2O_4$	25.5	31.9	36.4	39.9	43.8		53.2		63.6	69.2	75.3
Potassium perchlorate	$KClO_4$	0.76	1.06	1.68	2.56	3.73		7.3		13.4	17.7	22.3
Potassium periodate	KIO_4	0.17	0.28	0.42	0.65	1		2.1		4.4	5.9	
Potassium permanganate	$KMnO_4$	2.83	4.31	6.34	9.03	12.6	16.9	22.1				
Potassium persulfate	$K_2S_2O_8$			4.7								
Potassium phosphate	K_3PO_4		81.5	92.3	108	133						
Potassium selenate	K_2SeO_4	107	109	111	113	115		119		121		122
Potassium sulfate	K_2SO_4	7.4	9.3	11.1	13	14.8		18.2		21.4	22.9	24.1
Potassium tetraphenylborate	$KB(C_6H_5)_4$			1.8×10^{-5}								
Potassium thiocyanate	KSCN	177	198	224	255	289		372		492	571	675
Potassium thiosulfate	$K_2S_2O_3$	96		155	175	205		238		293	312	
Potassium tungstate	K_2WO_4			51.5								
Praseodymium(III) acetate	$Pr(C_2H_3O_2)_3{\cdot}H_2O$			32								
Praseodymium(III) bromate	$Pr(BrO_3)_3$	55.9	73	91.8	114	144						
Praseodymium(III) chloride	$PrCl_3$			104								
Praseodymium(III) molybdate	$Pr_2(MoO_4)_3$			0.0015								
Praseodymium(III) nitrate	$Pr(NO_3)_3$			112	162	178						
Praseodymium(III) sulfate	$Pr_2(SO_4)_3$	19.8	15.6	12.6	9.89	2.56		5.04		3.5	1.1	0.91

R

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Radium chloride	$RaCl_2$			19.6								
Radium iodate	$Ra(IO_3)_2$			0.04								
Radium nitrate	$Ra(NO_3)_2$			12								
Radium sulfate	$RaSO_4$			2.1×10^{-4}								
Raffinose	$C_{18}H_{32}O_{16}{\cdot}5H_2O$			14								
Rubidium acetate	$RbC_2H_3O_2$					86						
Rubidium bromate	$RbBrO_3$				3.6	5.1						
Rubidium bromide	RbBr	90	99	108	119	132		158				
Rubidium chlorate	$RbClO_3$	2.1	3.1	5.4	8	11.6		22		38	49	63
Rubidium chloride	RbCl	77	84	91	98	104		115		127	133	143
Rubidium chromate	Rb_2CrO_4	62	67.5	73.6	78.9	85.6		95.7				
Rubidium dichromate	$Rb_2Cr_2O_7$			5.9	10	15.2		32.3				
Rubidium fluoride	RbF			130.6 **(18 °C)**								
Rubidium fluorosilicate	Rb_2SiF_6			0.157								
Rubidium formate	$RbHCO_2$		443	554	614	694		900				
Rubidium hydrogen carbonate	$RbHCO_3$			110								
Rubidium hydroxide	RbOH			180								
Rubidium iodate	$RbIO_3$			1.96								
Rubidium iodide	RbI			144								
Rubidium nitrate	$RbNO_3$	19.5	33	52.9	81.2	117		200		310	374	452

Rubidium perchlorate	$RbClO_4$	1.09	1.19	1.55	2.2	3.26		6.27		11	15.5	22
Rubidium periodate	$RbIO_4$			0.648								
Rubidium permanganate	$RbMnO_4$	0.41		1.12 **(19 °C)**		2.34	3.25	4.68				
Rubidium phosphate	Rb_3PO_4			220								
Rubidium selenate	Rb_2SeO_4			159								
Rubidium sulfate	Rb_2SO_4	37.5	42.6	48.1	53.6	58.5		67.5		75.1	78.6	81.8

S

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Samarium acetate	$Sm(C_2H_3O_2)_3{\cdot}3H_2O$			15								
Samarium bromate	$Sm(BrO_3)_3$	34.2	47.6	62.5	79	98						
Samarium chloride	$SmCl_3$		92.4	93.4	94.6	96.9						
Samarium sulfate	$Sm_2(SO_4)_3{\cdot}8H_2O$			2.7	3.1							
Scandium oxalate	$Sc_2(C_2O_4)_3{\cdot}6H_2O$			0.006								
Scandium sulfate	$Sc_2(SO_4)_3{\cdot}5H_2O$			54.6								
Silicon dioxide	SiO_2			0.012								
Silver acetate	$AgC_2H_3O_2$	0.73	0.89	1.05	1.23	1.43		1.93		2.59		
Silver azide	AgN_3			7.931×10^{-4}								
Silver bromate	$AgBrO_3$		0.11	0.16	0.23	0.32		0.57		0.94	1.33	
Silver bromide	$AgBr$			1.328×10^{-5}								
Silver carbonate	Ag_2CO_3			0.003489								
Silver chlorate	$AgClO_3$		10.4	15.3	20.9	26.8						
Silver chloride	$AgCl$			1.923×10^{-4}			0.00052					
Silver chlorite	$AgClO_2$			0.248								
Silver chromate	Ag_2CrO_4			0.002157								
Silver cyanide	$AgCN$			1.467×10^{-7}								
Silver dichromate	$Ag_2Cr_2O_7$			0.159								
Silver fluoride	AgF	85.9	120	172	190	203						
Silver nitrate	$AgNO_3$	122	167	216	265	311		440		585	652	733
Silver oxalate	$Ag_2C_2O_4$			0.00327								
Silver oxide	Ag_2O			0.0012								
Silver perchlorate	$AgClO_4$	455	484	525	594	635						793
Silver permanganate	$AgMnO_4$			0.9								
Silver sulfate	Ag_2SO_4	0.57	0.7	0.8	0.89	0.98		1.15		1.3	1.36	1.41
Silver vanadate	$AgVO_3$			0.01462								
Sodium acetate	$NaC_2H_3O_2$	36.2	40.8	46.4	54.6	65.6		139		153	161	170
Sodium azide	NaN_3	38.9	39.9	40.8								
Sodium benzoate	$NaC_7H_5O_2$			66								
Sodium borohydride	$NaBH_4$	25		55		88.5						
Sodium bromate	$NaBrO_3$	24.2	30.3	36.4	42.6	48.8		62.6		75.7		90.8
Sodium bromide	$NaBr$	80.2	85.2	90.8	98.4	107		118		120	121	121
Sodium carbonate	Na_2CO_3	7	12.5	21.5	39.7	49		46		43.9	43.9	45.5
Sodium chlorate	$NaClO_3$	79.6	87.6	95.9	105	115		137		167	184	204
Sodium chloride	$NaCl$	35.65	35.72	35.89	36.09	36.37	36.69	37.04	37.46	37.93	38.47	38.99
Sodium chromate	Na_2CrO_4	31.7	50.1	84	88	96		115		125		126

Sodium cyanide	NaCN	40.8	48.1	58.7	71.2	dec						
Sodium dichromate	$Na_2Cr_2O_7$	163	172	183	198	215		269		376	405	415
Monosodium phosphate	NaH_2PO_4	56.5	69.8	86.9	107	133		172		211	234	
Sodium fluoride	NaF	3.66		4.06	4.22	4.4		4.68		4.89		5.08
Sodium formate	HCOONa	43.9	62.5	81.2	102	108		122		138	147	160
Sodium hydrogen carbonate	$NaHCO_3$	7	8.1	9.6	11.1	12.7		16				
Sodium hydroxide	NaOH		98	109	119	129		174				
Sodium iodate	$NaIO_3$	2.48	4.59	8.08	10.7	13.3		19.8		26.6	29.5	33
Sodium iodide	NaI	159	167	178	191	205		257		295		302
Sodium metabisulfite	$Na_2S_2O_5$	45.1		65.3						88.7		96.3
Sodium metaborate	$NaBO_2$	16.4	20.8	25.4	31.4	40.4		63.9		84.5		125.2
Sodium molybdate	Na_2MoO_4	44.1	64.7	65.3	66.9	68.6		71.8				
Sodium nitrate	$NaNO_3$	73	80.8	87.6	94.9	102		122		148		180
Sodium nitrite	$NaNO_2$	71.2	75.1	80.8	87.6	94.9		111		133		160
Sodium oxalate	$Na_2C_2O_4$	2.69	3.05	3.41	3.81	4.18		4.93		5.71		6.5
Sodium perchlorate	$NaClO_4$	167	183	201	222	245		288		306		329
Sodium periodate	$NaIO_4$	1.83	5.6	10.3	19.9	30.4						
Sodium permanganate	$NaMnO_4$			90								
Sodium phosphate	Na_3PO_4	4.5	8.2	12.1	16.3	20.2		20.9		60	68.1	77
Sodium pyrophosphate	$Na_4P_2O_7$	2.26										
Sodium selenate	Na_2SeO_4	13.3	25.2	26.9	77	81.8		78.6		74.8	73	72.7
Sodium sulfate	Na_2SO_4	4.9	9.1	19.5	40.8	48.8		45.3		43.7	42.7	42.5
Sodium sulfite	Na_2SO_3			27.0								
Sodium tetraborate (decahydrate)	$Na_2B_4O_7{\cdot}10H_2O$	2	2.3	2.5	4	6	10	15				
Sodium tetraborate (pentahydrate)	$Na_2B_4O_7{\cdot}5H_2O$								20	23	28	35
Sodium tetraborate (tetrahydrate)	$Na_2B_4O_7{\cdot}4H_2O$								17	20	23	28
Sodium tetrafluoroborate	$NaBF_4$	72	85									210
Sodium tetraphenylborate	$NaB(C_6H_5)_4$			47								
Sodium thiosulfate	$Na_2S_2O_3$	71.5		73		77.6				90.8		97.2
Strontium acetate	$Sr(C_2H_3O_2)_2$	37	42.9	41.1	39.5	38.3		36.8		36.1	36.2	36.4
Strontium bromate	$Sr(BrO_3)_2{\cdot}H_2O$			30.9								41
Strontium bromide	$SrBr_2$	85.2	93.4	102	112	123		150		182		223
Strontium carbonate	$SrCO_3$			0.0011								0.065
Strontium chlorate	$Sr(ClO_3)_2$			175								
Strontium chloride	$SrCl_2$	43.5	47.7	52.9	58.7	65.3		81.8		90.5		101
Strontium chromate	$SrCrO_4$			0.085	0.090							
Strontium fluoride	SrF_2			1.2×10^{-4}								
Strontium formate	$Sr(HCO_2)_2$	9.1	10.6	12.7	15.2	17.8		25		31.9	32.9	34.4
Strontium hydroxide	$Sr(OH)_2$	0.41[3]				1.77						21.83
Strontium iodate	$Sr(IO_3)_2$			0.19								0.35
Strontium iodide	SrI_2	165		178		192		218		270	365	383
Strontium molybdate	$SrMoO_4$			0.01107								

Strontium nitrate	$Sr(NO_3)_2$	39.5	54.9	70.8	87.6	91.3	92.6	94.0	97.2	99.0	101.1	
Strontium perchlorate[4]	$Sr(ClO_4)_2$	233.8	258.7	291.7	327.5	363.9						
Strontium selenate	$SrSeO_4$			0.656								
Strontium sulfate	$SrSO_4$	0.0113	0.0129	0.0132	0.0138	0.0141		0.0131		0.0116	0.0115	
Strontium thiosulfate	$SrS_2O_3·5H_2O$		2.5									
Strontium tungstate	$SrWO_4$			$3.957×10^{-4}$								
Sucrose	$C_{12}H_{22}O_{11}$	181.9	190.6	201.9	216.7	235.6	259.6	288.8	323.7	365.1	414.9	476.0
Sulfur dioxide	SO_2			9.4								

T

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Terbium bromate	$Tb(BrO_3)_3·9H_2O$	66.4	89.7	117	152	198						
Terbium sulfate	$Tb_2(SO_4)_3·8H_2O$			3.56								
Thallium(I) azide	TlN_3	0.171	0.236	0.364								
Thallium(I) bromate	$TlBrO_3$			0.306								
Thallium(I) bromide	TlBr	0.0238	0.032	0.0476	0.068	0.097		0.204				
Thallium(I) carbonate	Tl_2CO_3			5.3								
Thallium(I) chlorate	$TlClO_3$	2		3.92		12.7				36.6		57.3
Thallium(I) chloride	TlCl	0.17	0.24	0.34	0.46	0.60	0.80	1.02		1.60		2.41
Thallium(I) cyanide	TlCN			16.8								
Thallium(I) fluoride	TlF		78.6 **(15 °C)**									
Thallium(I) formate	$TlHCO_2$			500								
Thallium(I) hydroxide	TlOH	25.4	29.6	35	40.4	49.4		73.3		106	126	150
Thallium(I) iodate	$TlIO_3$			0.06678								
Thallium(I) iodide	TlI	0.002		0.006		0.015		0.035		0.07		0.12
Thallium(I) nitrate	$TlNO_3$	3.9	6.22	9.55	14.3	21		46.1		110	200	414
Thallium(I) oxalate	$Tl_2C_2O_4$			1.83								
Thallium(I) perchlorate	$TlClO_4$	6	8.04	13.1	19.7	28.3		50.8		81.5		
Thallium(I) phosphate	Tl_3PO_4			0.15								
Thallium(I) pyrophosphate	$Tl_4P_2O_7$			40								
Thallium(I) selenate	Tl_2SeO_4		2.17	2.8						8.5		10.8
Thallium(I) sulfate	Tl_2SO_4	2.73	3.7	4.87	6.16	7.53		11		14.6	16.5	18.4
Thallium(I) thiocyanate[5]	TlSCN		$4.6×10^{-5}$	$1.32×10^{-4}$ **(21 °C)**	$2.89×10^{-4}$	$6.76×10^{-4}$						
Thallium(I) vanadate	$TlVO_3$			0.87								
Thorium(IV) fluoride	$ThF_4·4H_2O$			0.914								
Thorium(IV) iodate	$Th(IO_3)_4$			0.03691								
Thorium(IV) nitrate	$Th(NO_3)_4$	186	187	191								
Thorium(IV) selenate	$Th(SeO_4)_2·9H_2O$	0.65										
Thorium(IV) sulfate	$Th(SO_4)_2·9H_2O$	0.74	0.99	1.38	1.99	3						
Thulium(III) nitrate	$Tm(NO_3)_3$			212								
Tin(II) bromide	$SnBr_2$	85										
Tin(II) chloride	$SnCl_2$	84										
Tin(II) fluoride	SnF_2			30								

Tin(II) iodide	SnI_2			0.99	1.17	1.42		2.11		3.04	3.58	4.2
Tin(II) sulfate	$SnSO_4$			18.9								
Trehalose	$C_{12}H_{22}O_{11}$			68.9								

U, V, and X

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Uranyl acetate	$UO_2(C_2H_3O_2)_2{\cdot}2H_2O$			7.69								
Uranyl chloride	UO_2Cl_2			320								
Uranyl formate	$UO_2(HCO_2)_2{\cdot}H_2O$			7.2								
Uranyl iodate	$UO_2(IO_3)_2{\cdot}H_2O$			0.124								
Uranyl nitrate	$UO_2(NO_3)_2$	98	107	122	141	167		317		388	426	474
Uranyl oxalate	$UO_2C_2O_4$		0.45	0.5	0.61	0.8		1.22		1.94		3.16
Uranyl sulfate	$UO_2SO_4{\cdot}3H_2O$			21								
Urea	$CO(NH_2)_2$	66.7		108		167		251		400		733
Vanadium(V) oxide	V_2O_5			0.8								
Xenon	Xe	24.1 ml		11.9 ml			8.4 ml			7.12 ml		
Xylose	$C_5H_{10}O_5$			117								

Y

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Ytterbium(III) nitrate	$Yb(NO_3)_3$			239								
Ytterbium(III) sulfate	$Yb_2(SO_4)_3$	44.2	37.5	38.4	22.2	17.2		10.4		6.4	5.8	4.7
Yttrium(III) acetate	$Y(C_2H_3O_2)_3{\cdot}4H_2O$			9.03								
Yttrium(III) bromate	$Y(BrO_3)_3{\cdot}9H_2O$			168								
Yttrium(III) bromide	YBr_3	63.9		75.1		87.3		101		116	123	
Yttrium(III) chloride	YCl_3	77.3	78.1	78.8	79.6	80.8						
Yttrium(III) fluoride	YF_3			0.005769								
Yttrium(III) nitrate	$Y(NO_3)_3$	93.1	106	123	143	163		200				
Yttrium(III) sulfate	$Y_2(SO_4)_3$	8.05	7.67	7.3	6.78	6.09		4.44		2.89	2.2	

Z

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Zinc acetate	$Zn(C_2H_3O_2)_2$			30								
Zinc bromide	$ZnBr_2$	389		446	528	591		618		645		672
Zinc carbonate	$ZnCO_3$			4.692×10^{-5}								
Zinc chlorate	$Zn(ClO_3)_2$	145	152	200	209	223						
Zinc chloride	$ZnCl_2$	342	363	395	437	452		488		541		614
Zinc cyanide	$Zn(CN)_2$			0.058								
Zinc fluoride	ZnF_2			1.6								
Zinc formate	$Zn(HCO_2)_2$	3.7	4.3	5.2	6.1	7.4		11.8		21.2	28.8	
Zinc iodate	$Zn(IO_3)_2{\cdot}2H_2O$			0.07749								
Zinc iodide	ZnI_2	430		432		445		467		490		510
Zinc nitrate	$Zn(NO_3)_2$	98			138	211						
Zinc oxalate	$ZnC_2O_4{\cdot}2H_2O$			1.38×10^{-9}								

Zinc oxide	ZnO			4.20×10^{-4}								
Zinc permanganate	$Zn(MnO_4)_2$			33.3								
Zinc sulfate	$ZnSO_4$	41.6	47.2	53.8	61.3	70.5		75.4		71.1		60.5
Zinc sulfite	$ZnSO_3{\cdot}2H_2O$			0.16								
Zinc tartrate	$ZnC_4H_4O_6$			0.022	0.041	0.06		0.104		0.59		
Zirconium fluoride	ZrF_4			1.32								
Zirconium sulfate	$Zr(SO_4)_2{\cdot}4H_2O$			52.5								

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- Chemicalc v4.0 - software that includes data on solubility
- Learning, Food resources (https://web.archive.org/web/20061207132126/http://food.oregonstate.edu/learn/sugar.html)
- Kaye and Laby Online (http://www.kayelaby.npl.co.uk/)
- ChemBioFinder.com (https://web.archive.org/web/20141109053519/http://www.chemfinder.com/)(registration required)

External links

- Solubility Database (http://srdata.nist.gov/solubility/index.aspx) - International Union of Pure and Applied Chemistry / National Institute of Standards and Technology
- CRC Handbook of Chemistry and Physics (http://hbcponline.com) - Online resource that includes solubility data (requires subscription)

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