

# Solubility table

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See also: [Solubility chart](#)

The table below provides information on the variation of [solubility](#) of different substances (mostly [inorganic compounds](#)) in water with [temperature](#), at 1 [atmosphere pressure](#). Units of solubility are given in [grams](#) per 100 grams of water (g/100g), unless shown otherwise. The substances are listed in alphabetical order.

In general, substances have to be exposed under boiling point for a short while to fully dissolve.

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## A [edit]

Substance	Formula	0 °C	10 °C	15 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<a href="#">Actinium(III) hydroxide</a>	<a href="#">Ac(OH)<sub>3</sub></a>				0.0021								
<a href="#">Aluminium chloride</a>	<a href="#">AlCl<sub>3</sub></a>	43.9	44.9		45.8	46.6	47.3		48.1		48.6		49.0
<a href="#">Aluminium fluoride</a>	<a href="#">AlF<sub>3</sub></a>	0.56	0.56		0.67	0.78	0.91		1.1		1.32		1.72
<a href="#">Aluminium hydroxide</a>	<a href="#">Al(OH)<sub>3</sub></a>				0.0001								
<a href="#">Aluminium nitrate</a>	<a href="#">Al(NO<sub>3</sub>)<sub>3</sub></a>	60	66.7		73.9	81.8	88.7	96.0	106	120	132	153	160
<a href="#">Aluminium perchlorate</a>	<a href="#">Al(ClO<sub>4</sub>)<sub>3</sub></a>	122	128		133								
<a href="#">Aluminium sulfate</a>	<a href="#">Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></a>	31.2	33.5		36.4	40.4	45.8	52.2	59.2	66.2	73	80.8	89.0
<a href="#">Ammonia (mL/mL)</a>	<a href="#">NH<sub>3</sub></a>	1176	900		702	565	428	333	252	188	138	100	88
<a href="#">Ammonium acetate</a>	<a href="#">NH<sub>4</sub>C<sub>2</sub>H<sub>3</sub>O<sub>2</sub></a>	102			143		204		311		533		
<a href="#">Ammonium azide</a>	<a href="#">NH<sub>4</sub>N<sub>3</sub></a>	16			25.3		37.1						
<a href="#">Ammonium benzoate</a>	<a href="#">NH<sub>4</sub>C<sub>7</sub>H<sub>5</sub>O<sub>2</sub></a>		19.6		21.3								83
<a href="#">Ammonium bicarbonate</a>	<a href="#">NH<sub>4</sub>HCO<sub>3</sub></a>	11.9	16.1		21.7	28.4	36.6		59.2		109	<a href="#">dec</a>	
<a href="#">Ammonium bromide</a>	<a href="#">NH<sub>4</sub>Br</a>	60.6	68.1		76.4	83.2	91.2	99.2	108	117	125	135	145
<a href="#">Ammonium carbonate</a>	<a href="#">(NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub>·H<sub>2</sub>O</a>	55.8			10				<a href="#">dec</a>				
<a href="#">Ammonium chlorate</a>	<a href="#">NH<sub>4</sub>ClO<sub>3</sub></a>				28.7								
<a href="#">Ammonium chloride</a>	<a href="#">NH<sub>4</sub>Cl</a>	29.4	33.2		37.2	41.4	45.8	50.4	55.3	60.2	65.6	71.2	77.3
<a href="#">Ammonium hexachloroplatinate</a>	<a href="#">(NH<sub>4</sub>)<sub>2</sub>PtCl<sub>6</sub></a>	0.289	0.374		0.499	0.637	0.815		1.44		2.16	2.61	3.36
<a href="#">Ammonium chromate</a>	<a href="#">(NH<sub>4</sub>)<sub>2</sub>CrO<sub>4</sub></a>	25	29.2		34	39.3	45.3	51.9	59.0	71.2	76.1		
<a href="#">Ammonium dichromate</a>	<a href="#">(NH<sub>4</sub>)<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub></a>	18.2	25.5		35.6	46.5	58.5	71.4	86.0		115		156
<a href="#">Ammonium dihydrogen arsenate</a>	<a href="#">NH<sub>4</sub>H<sub>2</sub>AsO<sub>4</sub></a>	33.7			48.7		63.8		83		107	122	
<a href="#">Ammonium dihydrogen phosphate</a>	<a href="#">NH<sub>4</sub>H<sub>2</sub>PO<sub>4</sub></a>	22.7	39.5		37.4	46.4	56.7	69.0	82.5	98.6	118.3	142.8	173.2
<a href="#">Ammonium fluoride</a>	<a href="#">NH<sub>4</sub>F</a>	100											

Ammonium fluorosilicate	$(\text{NH}_4)_2\text{SiF}_6$	12.28	16.41		18.6	25.0	31.6	35.4	40.4	44.9	48.1(75 )		61.0
Ammonium formate	$\text{NH}_4\text{HCO}_2$	102			143		204		311		533		
Ammonium hydrogen phosphate	$(\text{NH}_4)_2\text{HPO}_4$	42.9	62.9		68.9	75.1	81.8	89.2	97.2	106	110	112	121
Ammonium hydrogen sulfate	$\text{NH}_4\text{HSO}_4$				100								
Ammonium hydrogen tartrate	$\text{NH}_4\text{HC}_4\text{H}_4\text{O}_6$		1.88		2.7								
Ammonium iodate	$\text{NH}_4\text{IO}_3$			2.6									14.5
Ammonium iodide	$\text{NH}_4\text{I}$	155	163		172	182	191	200	209	219	229		250
Ammonium nitrate	$\text{NH}_4\text{NO}_3$	118	150		192	242	297	344	421	499	580	740	871
Ammonium orthoperiodate	$(\text{NH}_4)_5\text{IO}_6$				2.7								
Ammonium oxalate	$(\text{NH}_4)_2\text{C}_2\text{O}_4$	2.2	3.21		4.45	6.09	8.18	10.3	14.0		22.4	27.9	34.7
Ammonium perchlorate	$\text{NH}_4\text{ClO}_4$	11.56	16.4		20.85		30.58		39.05		48.19		57.01
Ammonium permanganate	$\text{NH}_4\text{MnO}_4$			8.0					dec				
Ammonium perrhenate	$\text{NH}_4\text{ReO}_4$	2.8			6.2		12.0		20.7		32.3	39.1	
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$	9.40			20.3			37.7					
Ammonium selenate	$(\text{NH}_4)_2\text{SeO}_4$	96	105		115	126	143		192				
Ammonium sulfate	$(\text{NH}_4)_2\text{SO}_4$	70.6	73		75.4	78.1	81.2	84.3	87.4		94.1		103
Ammonium aluminium sulfate	$\text{NH}_4\text{Al}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$	2.4	5.0		7.4	10.5	14.6	19.6	26.7	37.7	53.9	98.2	121
Ammonium sulfite	$(\text{NH}_4)_2\text{SO}_3$	47.9	54		60.8	68.8	78.4		104		144	150	153
Ammonium tartrate	$(\text{NH}_4)_2\text{C}_4\text{H}_4\text{O}_6$	45	55		63	70.5	76.5		86.9				
Ammonium thiocyanate	$\text{NH}_4\text{SCN}$	120	144		170	208	234	235	346				
Ammonium thiosulfate	$(\text{NH}_4)_2\text{S}_2\text{O}_3$				173		205				269		
Ammonium vanadate	$\text{NH}_4\text{VO}_3$				0.48	0.84	1.32	1.78	2.42	3.05			7.0
Aniline	$\text{C}_6\text{H}_7\text{N}$				3.6								
Antimony trifluoride	$\text{SbF}_3$	385			444	562	dec						
Antimony sulfide	$\text{Sb}_2\text{S}_3$				$1.8 \times 10^{-4}$								
Antimony trichloride	$\text{SbCl}_3$	602			910	1090	1370	1917	4531	dec			
Argon (Unit:mL/mL)	Ar	0.056	0.0405		0.0336	0.0288	0.0252	0.0223					
Arsenic pentasulfide	$\text{As}_2\text{S}_5$	0.0014											
Arsenic pentoxide	$\text{As}_2\text{O}_5$	59.5	62.1		65.8	70.6	71.2		73.0		75.1		76.7
Arsenious sulfide	$\text{As}_2\text{S}_3$				0.0004								
Arsenic trioxide	$\text{As}_2\text{O}_3$	1.21	1.58		1.80		2.93	3.43	4.44	5.37	5.89	6.55	9
Arsine (Unit:mL/mL)	$\text{AsH}_3$				0.2								

**B** [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Barium acetate	$\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$	58.8	62	72	75	78.5	77	75	74	74		
Barium arsenate	$\text{Ba}_3(\text{AsO}_4)_2$			$2.586 \times 10^{-9}$								

Barium azide	$\text{Ba}(\text{N}_3)_2$	12.5	16.1	17.4					24.75			
Barium bromate monohydrate	$\text{Ba}(\text{BrO}_3)_2 \cdot \text{H}_2\text{O}$	0.29	0.44	0.65	0.95	1.31	1.75	2.27	3.01	3.65	4.45	5.71
Barium bromide	$\text{BaBr}_2$	98	101	104	109	114		123		135		149
Barium carbonate	$\text{BaCO}_3$			$1.409 \times 10^{-3}$								
Barium chlorate	$\text{Ba}(\text{ClO}_3)_2$	20.3	26.9	33.9	41.6	49.7		66.7		84.8		105
Barium chloride	$\text{BaCl}_2$	31.2	33.5	35.8	38.1	40.8		46.2		52.5	55.8	59.4
Barium chlorite	$\text{Ba}(\text{ClO}_2)_2$	43.9	44.6	45.4		47.9		53.8		66.6		80.8
Barium chromate	$\text{BaCrO}_4$			$2.775 \times 10^{-4}$								
Barium cyanide	$\text{Ba}(\text{CN})_2$			80								
Barium ferrocyanide	$\text{Ba}_2\text{Fe}(\text{CN})_6$			0.009732								
Barium fluoride	$\text{BaF}_2$		0.159	0.16	0.161							
Barium fluorosilicate	$\text{BaSiF}_6$			0.028								
Barium formate	$\text{Ba}(\text{HCO}_2)_2$	26.2	28	31.9	34		38.6		44.2	47.6	51.3	
Barium hydrogen phosphate	$\text{BaHPO}_4$			0.013								
Barium hydrogen phosphite	$\text{BaHPO}_3$			0.687								
Barium hydroxide	$\text{Ba}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$	1.67	2.48	3.89	5.59	8.22	11.7	20.9		101		
Barium iodate	$\text{Ba}(\text{IO}_3)_2$			0.035	0.046	0.057						0.2
Barium iodide	$\text{BaI}_2$	182	201	223	250			264			291	301
Barium molybdate	$\text{BaMoO}_4$			0.006								
Barium nitrate	$\text{Ba}(\text{NO}_3)_2$	4.95	6.77	9.02	11.5	14.1		20.4		27.2		34.4
Barium nitrite	$\text{Ba}(\text{NO}_2)_2$	50.3	60	72.8		102		151		222	261	325
Barium oxalate	$\text{BaC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$			0.003								
Barium oxide	$\text{BaO}$			3.48							90.8	
Barium perchlorate	$\text{Ba}(\text{ClO}_4)_2$	239		336		416		495		575		653
Barium permanganate	$\text{Ba}(\text{MnO}_4)_2$			0.015								
Barium pyrophosphate	$\text{Ba}_2\text{P}_2\text{O}_7$			0.009								
Barium selenate	$\text{BaSeO}_4$			0.005								
Barium sulfate	$\text{BaSO}_4$			$2.448 \times 10^{-4}$	$2.85 \times 10^{-4}$							
Barium sulfide	$\text{BaS}$	2.88	4.89	7.86	10.4	14.9		27.7		49.9	67.3	60.3
Beryllium carbonate	$\text{BeCO}_3$			0.218								
Beryllium chloride	$\text{BeCl}_2$		42	42								
Beryllium molybdate	$\text{BeMoO}_4$			3.02								
Beryllium nitrate	$\text{Be}(\text{NO}_3)_2$	97	102	108	113	125		178				
Beryllium oxalate	$\text{BeC}_2\text{O}_4 \cdot 3\text{H}_2\text{O}$			63.5								
Beryllium perchlorate	$\text{Be}(\text{ClO}_4)_2$			147								
Beryllium selenate	$\text{BeSeO}_4 \cdot 4\text{H}_2\text{O}$			49								
Beryllium sulfate	$\text{BeSO}_4 \cdot 4\text{H}_2\text{O}$	37	37.6	39.1	41.4	45.8		53.1		67.2		82.8
Bismuth arsenate	$\text{BiAsO}_4$			$7.298 \times 10^{-4}$								
Bismuth hydroxide	$\text{Bi}(\text{OH})_3$			$2.868 \times 10^{-7}$								
Bismuth iodide	$\text{BiI}_3$			$7.761 \times 10^{-4}$								
Bismuth phosphate	$\text{BiPO}_4$			$1.096 \times 10^{-10}$								
Bismuth sulfide	$\text{Bi}_2\text{S}_3$			$1.561 \times 10^{-20}$								
Boric acid	$\text{H}_3\text{BO}_3$	2.52	3.49	4.72	6.23	8.08	10.27	12.97	15.75	19.10	23.27	27.53
Boron trioxide	$\text{B}_2\text{O}_3$			2.2								
Bromine monochloride	$\text{BrCl}$			1.5								

C [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Cadmium arsenate	<span>Cd</span> <span><span><span>3</span></span></span> ( <span>AsO</span> <span><span><span>4</span></span></span> ) <span><span><span>2</span></span></span>			<span>7.091</span> <span>×</span> <span>10</span> <sup>-6</sup>								
Cadmium benzoate	<span>Cd</span> ( <span>C</span> <span><span><span>7</span></span></span> <span>H</span> <span><span><span>5</span></span></span> <span>O</span> <span><span><span>2</span></span></span> ) <span><span><span>2</span></span></span>			2.81								
Cadmium bromate	<span>Cd</span> ( <span>BrO</span> <span><span><span>3</span></span></span> ) <span><span><span>2</span></span></span>			125								
Cadmium bromide	<span>CdBr</span> <span><span><span>2</span></span></span>	56.3	75.4	98.8	129	152		153		156		160
Cadmium carbonate	<span>CdCO</span> <span><span><span>3</span></span></span>			<span>3.932</span> <span>×</span> <span>10</span> <sup>-5</sup>								
Cadmium chlorate	<span>Cd</span> ( <span>ClO</span> <span><span><span>3</span></span></span> ) <span><span><span>2</span></span></span>	299	308	322	348	376		455				
Cadmium chloride	<span>CdCl</span> <span><span><span>2</span></span></span>	100	135	135	135	135		136		140		147
Cadmium cyanide	<span>Cd</span> ( <span>CN</span> ) <span><span><span>2</span></span></span>			0.022								
Cadmium ferrocyanide	<span>Cd</span> <span><span><span>2</span></span></span> <span>Fe</span> ( <span>CN</span> ) <span><span><span>6</span></span></span>			<span>8.736</span> <span>×</span> <span>10</span> <sup>-5</sup>								
Cadmium fluoride	<span>CdF</span> <span><span><span>2</span></span></span>			4								
Cadmium formate	<span>Cd</span> ( <span>HCO</span> <span><span><span>2</span></span></span> ) <span><span><span>2</span></span></span>	8.3	11.1	14.4	18.6	25.3		59.5		80.5	85.2	94.6
Cadmium hydroxide	<span>Cd</span> ( <span>OH</span> ) <span><span><span>2</span></span></span>			<span>2.697</span> <span>×</span> <span>10</span> <sup>-4</sup>								
Cadmium iodate	<span>Cd</span> ( <span>IO</span> <span><span><span>3</span></span></span> ) <span><span><span>2</span></span></span>			0.097								
Cadmium iodide	<span>CdI</span> <span><span><span>2</span></span></span>	78.7		84.7	87.9	92.1		100		111		125
Cadmium nitrate	<span>Cd</span> ( <span>NO</span> <span><span><span>3</span></span></span> ) <span><span><span>2</span></span></span>	122		136	150	194		310		713		
Cadmium oxalate	<span>CdC</span> <span><span><span>2</span></span></span> <span>O</span> <span><span><span>4</span></span></span> <span>·</span> <span>3H</span> <span><span><span>2</span></span></span> <span>O</span>			0.006046								
Cadmium perchlorate	<span>Cd</span> ( <span>ClO</span> <span><span><span>4</span></span></span> ) <span><span><span>2</span></span></span>		180	188	195	203		221		243		272
Cadmium phosphate	<span>Cd</span> <span><span><span>3</span></span></span> ( <span>PO</span> <span><span><span>4</span></span></span> ) <span><span><span>2</span></span></span>			<span>6.235</span> <span>×</span> <span>10</span> <sup>-6</sup>								
Cadmium selenate	<span>CdSeO</span> <span><span><span>4</span></span></span>	72.5	68.4	64	58.9	55		44.2		32.5	27.2	22
Cadmium sulfate	<span>CdSO</span> <span><span><span>4</span></span></span>	75.4	76	76.6		78.5		81.8		66.7	63.1	60.8
Cadmium sulfide	<span>CdS</span>			<span>1.292</span> <span>×</span> <span>10</span> <sup>-12</sup>								
Cadmium tungstate	<span>CdWO</span> <span><span><span>4</span></span></span>			0.04642								
Caesium acetate	<span>CsC</span> <span><span><span>2</span></span></span> <span>H</span> <span><span><span>3</span></span></span> <span>O</span> <span><span><span>2</span></span></span>			1010								
Caesium azide	<span>CsN</span> <span><span><span>3</span></span></span>			307								
Caesium bromate	<span>CsBrO</span> <span><span><span>3</span></span></span>	0.21		3.66	4.53	5.3						
Caesium bromide	<span>CsBr</span>			108								
Caesium chlorate	<span>CsClO</span> <span><span><span>3</span></span></span>		3.8	6.2	9.5	13.8		26.2		45	58	79
Caesium chloride	<span>CsCl</span>	146	175	187	197	208		230		250	260	271
Caesium chromate	<span>Cs</span> <span><span><span>2</span></span></span> <span>CrO</span> <span><span><span>4</span></span></span>		71.4									
Caesium fluoride	<span>CsF</span>			322								
Caesium fluoroborate	<span>CsBF</span> <span><span><span>4</span></span></span>			0.818								
Caesium formate	<span>CsHCO</span> <span><span><span>2</span></span></span>	335	381	450	694							
Caesium iodate	<span>CsIO</span> <span><span><span>3</span></span></span>			2.6								
Caesium iodide	<span>CsI</span>	44.1	58.5	76.5	96	124		150		190	205	
Caesium nitrate	<span>CsNO</span> <span><span><span>3</span></span></span>	9.33	14.9	23	33.9	47.2		83.8		134	163	197
Caesium oxalate	<span>Cs</span> <span><span><span>2</span></span></span> <span>C</span> <span><span><span>2</span></span></span> <span>O</span> <span><span><span>4</span></span></span>			313								
Caesium perchlorate	<span>CsClO</span> <span><span><span>4</span></span></span>	0.8	1	1.6	2.6	4		7.3		14.4	20.5	30
Caesium permanganate	<span>CsMnO</span> <span><span><span>4</span></span></span>			0.228								
Caesium selenate	<span>Cs</span> <span><span><span>2</span></span></span> <span>SeO</span> <span><span><span>4</span></span></span>		244									
Caesium sulfate	<span>Cs</span> <span><span><span>2</span></span></span> <span>SO</span> <span><span><span>4</span></span></span>	167	173	179	184	190		200		210	215	200
Calcium acetate	<span>Ca</span> ( <span>C</span> <span><span><span>2</span></span></span> <span>H</span> <span><span><span>3</span></span></span> <span>O</span> <span><span><span>2</span></span></span> ) <span><span><span>2</span></span></span> <span>·</span> <span>2H</span> <span><span><span>2</span></span></span> <span>O</span>	37.4	36	34.7	33.8	33.2		32.7		33.5	31.1	29.7
Calcium arsenate	<span>Ca</span> <span><span><span>3</span></span></span> ( <span>AsO</span> <span><span><span>4</span></span></span> ) <span><span><span>2</span></span></span>			0.003629								
Calcium azide	<span>Ca</span> ( <span>N</span> <span><span><span>3</span></span></span> ) <span><span><span>2</span></span></span>			45								
Calcium benzoate	<span>Ca</span> ( <span>C</span> <span><span><span>7</span></span></span> <span>H</span> <span><span><span>5</span></span></span> <span>O</span> <span><span><span>2</span></span></span> ) <span><span><span>2</span></span></span> <span>·</span> <span>3H</span> <span><span><span>2</span></span></span> <span>O</span>	2.32	2.45	2.72	3.02	3.42		4.71		6.87	8.55	8.7

Calcium bicarbonate	$\text{Ca}(\text{HCO}_3)_2$	16.1		16.6		17.1		17.5	17.9		18.4
Calcium bromate	$\text{Ca}(\text{BrO}_3)_2$			230							
Calcium bromide	$\text{CaBr}_2$	125	132	143		213		278	295		312
Calcium carbonate (Aragonite)	$\text{CaCO}_3$ -Aragonite			$7.753 \times 10^{-4}$							
Calcium carbonate (Calcite)	$\text{CaCO}_3$ -Calcite			$6.17 \times 10^{-4}$							
Calcium chlorate	$\text{Ca}(\text{ClO}_3)_2$			209							
Calcium chloride	$\text{CaCl}_2$	59.5	64.7	74.5	100	128		137	147	154	159
Calcium chromate	$\text{CaCrO}_4$	4.5		2.25	1.83	1.49		0.83			
Calcium citrate	$\text{Ca}_3(\text{C}_6\text{H}_5\text{O}_7)_2$			0.095 (25 °C)							
Monocalcium phosphate	$\text{Ca}(\text{H}_2\text{PO}_4)_2$			1.8							
Calcium fluoride	$\text{CaF}_2$	0.008575									
Calcium fluorosilicate	$\text{CaSiF}_6$			0.518							
Calcium formate	$\text{Ca}(\text{HCO}_2)_2$	16.1		16.6		17.1		17.5	17.9		18.4
Dicalcium phosphate	$\text{CaHPO}_4$			0.004303							
Calcium hydroxide	$\text{Ca}(\text{OH})_2$	0.189	0.182	0.173	0.16	0.141		0.121	0.086	0.076	
Calcium iodate	$\text{Ca}(\text{IO}_3)_2$	0.09		0.24	0.38	0.52		0.65	0.66	0.67	0.67
Calcium iodide	$\text{CaI}_2$	64.6		66	67.6	70.8		74	78		81
Calcium molybdate	$\text{CaMoO}_4$			0.004099							
Calcium nitrate	$\text{Ca}(\text{NO}_3)_2$			121.2							
Calcium nitrate tetrahydrate	$\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$	102	115	129	152	191			358		363
Calcium nitrite	$\text{Ca}(\text{NO}_2)_2 \cdot 4\text{H}_2\text{O}$	63.9		84.5	104			134	151	166	178
Calcium oxalate	$\text{CaC}_2\text{O}_4$			$6.7 \times 10^{-4}$						0.0014	
Calcium oxide	$\text{CaO}$										5.7
Calcium perchlorate	$\text{Ca}(\text{ClO}_4)_2$			188							
Calcium permanganate	$\text{Ca}(\text{MnO}_4)_2$			338							
Calcium phosphate	$\text{Ca}_3(\text{PO}_4)_2$			0.002							
Calcium selenate	$\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$	9.73	9.77	9.22	8.79	7.14					
Calcium sulfate	$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$	0.223	0.244	0.255	0.264	0.265		0.244	0.234		0.205
Calcium tungstate	$\text{CaWO}_4$			0.002387							
Carbon dioxide	$\text{CO}_2$			0.1782							
Carbon monoxide	$\text{CO}$			0.0026							
Cerium(III) acetate	$\text{Ce}(\text{C}_2\text{H}_3\text{O}_2)_3$			0.35							
Cerium(III) chloride	$\text{CeCl}_3$			100							
Cerium(III) hydroxide	$\text{Ce}(\text{OH})_3$			$9.43 \times 10^{-5}$							
Cerium(III) iodate	$\text{Ce}(\text{IO}_3)_3$			0.123							
Cerium(III) nitrate	$\text{Ce}(\text{NO}_3)_3$			234							
Cerium(III) phosphate	$\text{CePO}_4$			$7.434 \times 10^{-11}$							
Cerium(III) selenate	$\text{Ce}_2(\text{SeO}_4)_3$	39.5	37.2	35.2	33.2	32.6		13.7	4.6		
Cerium(III) sulfate	$\text{Ce}_2(\text{SO}_4)_3 \cdot 2\text{H}_2\text{O}$	21.4		9.84	7.24	5.63		3.87			
Cerium(IV) hydroxide	$\text{Ce}(\text{OH})_4$			$1.981 \times 10^{-5}$							
Chromium(III) nitrate	$\text{Cr}(\text{NO}_3)_3$	108	124	130	152						
Chromium(III) perchlorate	$\text{Cr}(\text{ClO}_4)_3$	104	123	130							
Chromium(III) sulfate	$\text{Cr}_2(\text{SO}_4)_3 \cdot 18\text{H}_2\text{O}$			220							

Chromium(VI) oxide	$\text{CrO}_3$	61.7		63						67		
Cobalt(II) bromate	$\text{Co}(\text{BrO}_3)_2 \cdot 6\text{H}_2\text{O}$			45.5								
Cobalt(II) bromide	$\text{CoBr}_2$	91.9		112	128	163		227		241		257
Cobalt(II) chlorate	$\text{Co}(\text{ClO}_3)_2$	135	162	180	195	214		316				
Cobalt(II) chloride	$\text{CoCl}_2$	43.5	47.7	52.9	59.7	69.5		93.8		97.6	101	106
Cobalt(II) fluoride	$\text{CoF}_2$			1.36								
Cobalt(II) fluorosilicate	$\text{CoSiF}_6 \cdot 6\text{H}_2$			118								
Cobalt(II) iodate	$\text{Co}(\text{IO}_3)_2 \cdot 2\text{H}_2\text{O}$			1.02	0.9	0.88		0.82		0.73		0.7
Cobalt(II) iodide	$\text{CoI}_2$			203								
Cobalt(II) nitrate	$\text{Co}(\text{NO}_3)_2$	84	89.6	97.4	111	125		174		204	300	
Cobalt(II) nitrite	$\text{Co}(\text{NO}_2)_2$	0.076	0.24	0.4	0.61	0.85						
Cobalt oxalate	$\text{CoC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$			$2.6972 \times 10^{-9}$								
Cobalt(II) perchlorate	$\text{Co}(\text{ClO}_4)_2$			104								
Cobalt(II) sulfate	$\text{CoSO}_4$	25.5	30.5	36.1	42	48.8		55		53.8	45.3	38.9
Copper(I) chloride	$\text{CuCl}$			0.0099								
Copper(I) cyanide	$\text{CuCN}$			$1.602 \times 10^{-9}$								
Copper(I) hydroxide	$\text{CuOH}$			$8.055 \times 10^{-7}$								
Copper(I) iodide	$\text{CuI}$			$1.997 \times 10^{-5}$								
Copper(I) sulfide	$\text{Cu}_2\text{S}$			$1.361 \times 10^{-15}$								
Copper(I) thiocyanate	$\text{CuSCN}$			$8.427 \times 10^{-7}$								
Copper(II) bromide	$\text{CuBr}_2$	107	116	126	128	131						
Copper(II) carbonate	$\text{CuCO}_3$			$1.462 \times 10^{-4}$								
Copper(II) chlorate	$\text{Cu}(\text{ClO}_3)_2$			242								
Copper(II) chloride	$\text{CuCl}_2$	68.6	70.9	73	77.3	87.6		96.5		104	108	120
Copper(II) chromate	$\text{CuCrO}_4$			0.03407								
Copper(II) fluoride	$\text{CuF}_2$			0.075								
Copper(II) fluorosilicate	$\text{CuSiF}_6$	73.5	76.5	81.6	84.1	91.2				93.2		
Copper(II) formate	$\text{Cu}(\text{HCO}_2)_2$			12.5								
Copper(II) hydroxide	$\text{Cu}(\text{OH})_2$			$1.722 \times 10^{-6}$								
Copper(II) iodate	$\text{Cu}(\text{IO}_3)_2 \cdot 2\text{H}_2\text{O}$			0.109								
Copper(II) nitrate	$\text{Cu}(\text{NO}_3)_2$	83.5	100	125	156	163		182		208	222	247
Copper oxalate	$\text{CuC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$			$2.1627 \times 10^{-10}$								
Copper(II) perchlorate	$\text{Cu}(\text{ClO}_4)_2$				146							
Copper(II) selenate	$\text{CuSeO}_4$	12	14.5	17.5	21	25.2		36.5		53.7		
Copper(II) selenite	$\text{CuSeO}_3$			0.002761								
Copper(II) sulfate	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$	23.1	27.5	32	37.8	44.6		61.8		83.8		114
Copper(II) sulfide	$\text{CuS}$			$2.41 \times 10^{-17}$								

D [\[edit\]](#)

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Dysprosium(III) chromate	$\text{Dy}_2(\text{CrO}_4)_3 \cdot 10\text{H}_2\text{O}$			0.663								

E [\[edit\]](#)

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100
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												°C
Erbium(III) hydroxide	$\text{Er}(\text{OH})_3$			$1.363 \times 10^{-5}$								
Erbium(III) sulfate	$\text{Er}_2(\text{SO}_4)_3$			13.79								
Erbium(III) sulfate octahydrate	$\text{Er}_2(\text{SO}_4)_3 \cdot 8\text{H}_2\text{O}$			16.00		6.53						
Europium(III) hydroxide	$\text{Eu}(\text{OH})_3$			$1.538 \times 10^{-5}$								
Europium(III) sulfate	$\text{Eu}_2(\text{SO}_4)_3 \cdot 8\text{H}_2\text{O}$			2.56								

## F-G [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Ferrous ammonium sulfate	$(\text{NH}_4)_2\text{Fe}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$			26.9						73		
Fructose	$\text{C}_6\text{H}_{12}\text{O}_6$			375.0		538.0						
Gadolinium(III) acetate	$\text{Gd}(\text{C}_2\text{H}_3\text{O}_2)_3 \cdot 4\text{H}_2\text{O}$			11.6								
Gadolinium(III) bicarbonate	$\text{Gd}(\text{HCO}_3)_3$			5.61								
Gadolinium(III) bromate	$\text{Gd}(\text{BrO}_3)_3 \cdot 9\text{H}_2\text{O}$	50.2	70.1	95.6	126	166						
Gadolinium(III) hydroxide	$\text{Gd}(\text{OH})_3$			$1.882 \times 10^{-5}$								
Gadolinium(III) sulfate	$\text{Gd}_2(\text{SO}_4)_3$	3.98	3.3	2.6	2.32							
D-Galactose	$\text{C}_6\text{H}_{12}\text{O}_6$			10.3								68.3
Gallium hydroxide	$\text{Ga}(\text{OH})_3$			$8.616 \times 10^{-9}$								
Gallium oxalate	$\text{Ga}_2(\text{C}_2\text{O}_4)_3 \cdot 4\text{H}_2\text{O}$			0.4								
Gallium selenate	$\text{Ga}_2(\text{SeO}_4)_3 \cdot 16\text{H}_2\text{O}$			18.1								
Gallium hydroxide	$\text{Ga}(\text{OH})_3$			$8.616 \times 10^{-9}$								
D-Glucose	$\text{C}_6\text{H}_{12}\text{O}_6$			90								
Gold(III) chloride	$\text{AuCl}_3$			68								
Gold(V) oxalate	$\text{Au}_2(\text{C}_2\text{O}_4)_5$			0.258								

## H [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Hafnium(III) hydroxide	$\text{Hf}(\text{OH})_3$			$4.503 \times 10^{-4}$								
Hafnium(IV) hydroxide	$\text{Hf}(\text{OH})_4$			$4.503 \times 10^{-6}$								
Helium	He			0.6								
Holmium(III) hydroxide	$\text{Ho}(\text{OH})_3$			$2.519 \times 10^{-5}$								
Holmium(III) sulfate	$\text{Ho}_2(\text{SO}_4)_3 \cdot 8\text{H}_2\text{O}$			8.18	6.1	4.52						
Hydrogen chloride	HCl	81	75	70	65.5	61	57.5	53	50	47	43	40
Hydrogen sulfide	$\text{H}_2\text{S}$			0.33								

## I [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Indium(III) bromide	$\text{InBr}_3$			571								
Indium(III) chloride	$\text{InCl}_3$		210	212								
Indium(III) fluoride	$\text{InF}_3$			11.2								
Indium(III) hydroxide	$\text{In}(\text{OH})_3$			$3.645 \times 10^{-8}$								
Indium(III) iodate	$\text{In}(\text{IO}_3)_3$			0.067								
Indium(III) sulfide	$\text{In}_2\text{S}_3$			$2.867 \times 10^{-14}$								
Iron(II) bromide	$\text{FeBr}_2$	101	109	117	124	133		144		168	176	184

Iron(II) carbonate	$\text{FeCO}_3$			$6.554 \times 10^{-5}$								
Iron(II) chloride	$\text{FeCl}_2$	49.7	59	62.5	66.7	70		78.3		88.7	92.3	94.9
Iron(II) fluorosilicate	$\text{FeSiF}_6 \cdot 6\text{H}_2\text{O}$	72.1	74.4		77			84		88		100
Iron(II) hydroxide	$\text{Fe(OH)}_2$			$5.255 \times 10^{-5}$								
Iron(II) nitrate	$\text{Fe(NO}_3)_2 \cdot 6\text{H}_2\text{O}$	113	134									
Iron(II) oxalate	$\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$			0.008								
Iron(II) perchlorate	$\text{Fe(ClO}_4)_2 \cdot 6\text{H}_2\text{O}$			299								
Iron(II) sulfate	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$			28.8		40	48	60	73.3		101	79.9
Iron(III) arsenate	$\text{FeAsO}_4$			$1.47 \times 10^{-9}$								
Iron(III) chloride	$\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$	74.4		91.8	107							
Iron(III) fluoride	$\text{FeF}_3$			0.091								
Iron(III) hydroxide	$\text{Fe(OH)}_3$			$2.097 \times 10^{-9}$								
Iron(III) iodate	$\text{Fe(IO}_3)_3$			0.36								
Iron(III) nitrate	$\text{Fe(NO}_3)_3 \cdot 9\text{H}_2\text{O}$	112		138		175						
Iron(III) perchlorate	$\text{Fe(ClO}_4)_3$	289		368	422	478		772				
Iron(III) sulfate	$\text{Fe}_2(\text{SO}_4)_3 \cdot 9\text{H}_2\text{O}$											

L [edit]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Lactose	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$			8								
Lanthanum(III) acetate	$\text{La}(\text{C}_2\text{H}_3\text{O}_2)_3 \cdot \text{H}_2\text{O}$			16.9								
Lanthanum(III) bromate	$\text{La}(\text{BrO}_3)_3$	98	120	149	200							
Lanthanum(III) iodate	$\text{La}(\text{IO}_3)_3$			0.04575								
Lanthanum(III) molybdate	$\text{La}_2(\text{MoO}_4)_3$			0.002473								
Lanthanum(III) nitrate	$\text{La}(\text{NO}_3)_3$	100		136		168		247				
Lanthanum(III) selenate	$\text{La}_2(\text{SeO}_4)_3$	50.5	45	45	45	45		18.5		5.4	2.2	
Lanthanum(III) sulfate	$\text{La}_2(\text{SO}_4)_3$	3	2.72	2.33	1.9	1.67		1.26		0.91	0.79	0.68
Lanthanum(III) tungstate	$\text{La}_2(\text{WO}_4)_3 \cdot 3\text{H}_2\text{O}$			6.06								
Lead(II) acetate	$\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2$	19.8	29.5	44.3	69.8	116						
Lead(II) azide	$\text{Pb}(\text{N}_3)_2$			0.0249								
Lead(II) bromate	$\text{Pb}(\text{BrO}_3)_2$			7.92								
Lead(II) bromide	$\text{PbBr}_2$	0.45	0.63	0.973	1.12	1.5		2.29		3.32	3.86	4.55
Lead(II) carbonate	$\text{PbCO}_3$			$7.269 \times 10^{-5}$								
Lead(II) chlorate	$\text{Pb}(\text{ClO}_3)_2$			144								
Lead(II) chloride	$\text{PbCl}_2$	0.67	0.82	1.08	1.2	1.42		1.94		2.54	2.88	3.2
Lead(II) chromate	$\text{PbCrO}_4$			$1.71 \times 10^{-5}$								
Lead(II) ferrocyanide	$\text{PbFe}(\text{CN})_6$			$5.991 \times 10^{-4}$								
Lead(II) fluoride	$\text{PbF}_2$			0.0671								
Lead(II) fluorosilicate	$\text{PbSiF}_6$	190		222				403		428		463
Lead(II) hydrogen phosphate	$\text{PbHPO}_4$			$3.457 \times 10^{-4}$								
Lead(II) hydrogen phosphite	$\text{PbHPO}_3$			0.02187								
Lead(II) hydroxide	$\text{Pb(OH)}_2$			$1.615 \times 10^{-4}$								
Lead(II) iodate	$\text{Pb(IO}_3)_2$			0.0024								
Lead(II) iodide	$\text{PbI}_2$	0.044	0.056	0.0756	0.09	0.124		0.193		0.294		0.42
Lead(II) molybdate	$\text{PbMoO}_4$			$1.161 \times 10^{-5}$								



Lead(II) nitrate	$\text{Pb}(\text{NO}_3)_2$	37.5	46.2	54.3	63.4	72.1		91.6		111		133
Lead(II) oxalate	$\text{PbC}_2\text{O}_4$			$6.495 \times 10^{-4}$								
Lead(II) perchlorate	$\text{Pb}(\text{ClO}_4)_2 \cdot 3\text{H}_2\text{O}$			440								
Lead(II) selenate	$\text{PbSeO}_4$			0.0131								
Lead(II) sulfate	$\text{PbSO}_4$			0.00443								
Lead(II) sulfide	$\text{PbS}$			$6.767 \times 10^{-13}$								
Lead(II) tartrate	$\text{PbC}_4\text{H}_4\text{O}_6$			0.0025								
Lead(II) thiocyanate	$\text{Pb}(\text{SCN})_2$			0.553								
Lead(II) thiosulfate	$\text{PbS}_2\text{O}_3$			0.0202								
Lead(II) tungstate	$\text{PbWO}_4$			0.02838								
Lead(IV) hydroxide	$\text{Pb}(\text{OH})_4$			$7.229 \times 10^{-11}$								
Lithium acetate	$\text{LiC}_2\text{H}_3\text{O}_2$	31.2	35.1	40.8	50.6	68.6						
Lithium azide	$\text{LiN}_3$	61.3	64.2	67.2	71.2	75.4		86.6				100
Lithium benzoate	$\text{LiC}_7\text{H}_5\text{O}_2$	38.9	41.6	44.7	53.8							
Lithium bicarbonate	$\text{LiHCO}_3$			5.74								
Lithium bromate	$\text{LiBrO}_3$	154	166	179	198	221		269		308	329	355
Lithium bromide	$\text{LiBr}$	143	147	160	183	211		223		245		266
Lithium carbonate	$\text{Li}_2\text{CO}_3$	1.54	1.43	1.33	1.26	1.17		1.01		0.85		0.72
Lithium chlorate	$\text{LiClO}_3$	241	283	372	488	604		777				
Lithium chloride	$\text{LiCl}$	69.2	74.5	83.5	86.2	89.8		98.4		112	121	128
Lithium chromate	$\text{Li}_2\text{CrO}_4 \cdot 2\text{H}_2\text{O}$			142								
Lithium dichromate	$\text{Li}_2\text{Cr}_2\text{O}_7 \cdot 2\text{H}_2\text{O}$				151							
Lithium dihydrogen phosphate	$\text{LiH}_2\text{PO}_4$	126										
Lithium fluoride	$\text{LiF}$			0.27	0.135							
Lithium fluorosilicate	$\text{Li}_2\text{SiF}_6 \cdot 2\text{H}_2\text{O}$			73								
Lithium formate	$\text{LiHCO}_2$	32.3	35.7	39.3	44.1	49.5		64.7		92.7	116	138
Lithium hydrogen phosphite	$\text{Li}_2\text{HPO}_3$	4.43			9.97	7.61		7.11				6.03
Lithium hydroxide	$\text{LiOH}$	12.7	12.7	12.8	12.9	13.0	13.3	13.8		15.3		17.5
Lithium iodide	$\text{LiI}$	151	157	165	171	179		202		435	440	481
Lithium molybdate	$\text{Li}_2\text{MoO}_4$	82.6		79.5	79.5	78						73.9
Lithium nitrate	$\text{LiNO}_3$	53.4	60.8	70.1	138	152		175				
Lithium nitrite	$\text{LiNO}_2$	70.9	82.5	96.8	114	133		177		233	272	324
Lithium oxalate	$\text{Li}_2\text{C}_2\text{O}_4$			8								
Lithium perchlorate	$\text{LiClO}_4$	42.7	49	56.1	63.6	72.3		92.3		128	151	
Lithium permanganate	$\text{LiMnO}_4$			71.4								
Lithium phosphate	$\text{Li}_3\text{PO}_4$			0.039								
Lithium selenide	$\text{Li}_2\text{Se}$			57.7								
Lithium selenite	$\text{Li}_2\text{SeO}_3$	25	23.3	21.5	19.6	17.9		14.7		11.9	11.1	9.9
Lithium sulfate	$\text{Li}_2\text{SO}_4$	36.1	35.5	34.8	34.2	33.7		32.6		31.4	30.9	
Lithium tartrate	$\text{Li}_2\text{C}_4\text{H}_4\text{O}_6$	42	31.8	27.1	26.6	27.2		29.5				
Lithium thiocyanate	$\text{LiSCN}$			114	131	153						
Lithium vanadate	$\text{LiVO}_3$	2.5		4.82	6.28	4.38		2.67				
Lutetium(III) hydroxide	$\text{Lu}(\text{OH})_3$			$1.164 \times 10^{-5}$								
Lutetium(III) sulfate	$\text{Lu}_2(\text{SO}_4)_3 \cdot 8\text{H}_2\text{O}$			57.9								

**M** [\[edit\]](#)

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Magnesium acetate	$\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$	56.7	59.7	53.4	68.6	75.7		118				

Magnesium benzoate	$\text{Mg}(\text{C}_7\text{H}_5\text{O}_2)_2 \cdot \text{H}_2\text{O}$					5						
Magnesium bromate	$\text{Mg}(\text{BrO}_3)_2 \cdot 6\text{H}_2\text{O}$					58						
Magnesium bromide	$\text{MgBr}_2$	98	99	101	104	106		112				125
Magnesium carbonate	$\text{MgCO}_3$			0.039								
Magnesium chlorate	$\text{Mg}(\text{ClO}_3)_2$	114	123	135	155	178		242			268	
Magnesium chloride	$\text{MgCl}_2$	52.9	53.6	54.6	55.8	57.5		61		66.1	69.5	73.3
Magnesium chromate	$\text{MgCrO}_4 \cdot 7\text{H}_2\text{O}$			137								
Magnesium fluoride	$\text{MgF}_2$			0.007325								
Magnesium fluorosilicate	$\text{MgSiF}_6$	26.3		30.8		34.9		44.4				
Magnesium formate	$\text{Mg}(\text{HCO}_2)_2$	14	14.2	14.4	14.9	15.9		17.9		20.5	22.2	22.9
Magnesium hydroxide	$\text{Mg}(\text{OH})_2$			$9.628 \times 10^{-4}$								0.004
Magnesium iodate	$\text{Mg}(\text{IO}_3)_2$		7.2	8.6	10	11.7		15.2		15.5	15.6	
Magnesium iodide	$\text{MgI}_2$	120		140		173				186		
Magnesium molybdate	$\text{MgMoO}_4$			13.7								
Magnesium nitrate	$\text{Mg}(\text{NO}_3)_2$	62.1	66	69.5	73.6	78.9		78.9		91.6	106	
Magnesium oxalate	$\text{MgC}_2\text{O}_4$			0.104								
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$			49.6								
Magnesium phosphate	$\text{Mg}_3(\text{PO}_4)_2$			$2.588 \times 10^{-4}$								
Magnesium selenate	$\text{MgSeO}_4$	20	30.4	38.3	44.3	48.6		55.8				
Magnesium selenite	$\text{MgSeO}_3$			0.05454								
Magnesium sulfate	$\text{MgSO}_4$	25.5	30.4	35.1	39.7	44.7	50.4	54.8	59.2	54.8	52.9	50.2
Magnesium thiosulfate	$\text{MgS}_2\text{O}_3$			50								
Maltose	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$			108								
D-Mannose	$\text{C}_6\text{H}_{12}\text{O}_6$			248								
Manganese(II) bromide	$\text{MnBr}_2$	127	136	147	157	169		197		225	226	228
Manganese(II) carbonate	$\text{MnCO}_3$			$4.877 \times 10^{-5}$								
Manganese(II) chloride	$\text{MnCl}_2$	63.4	68.1	73.9	80.8	88.5		109		113	114	115
Manganese(II) ferrocyanide	$\text{Mn}_2\text{Fe}(\text{CN})_6$			0.001882								
Manganese(II) fluoride	$\text{MnF}_2$			0.96		0.67		0.44				0.48
Manganese(II) fluorosilicate	$\text{MnSiF}_6 \cdot 6\text{H}_2\text{O}$			140								
Manganese(II) hydroxide	$\text{Mn}(\text{OH})_2$			$3.221 \times 10^{-4}$								
Manganese(II) nitrate	$\text{Mn}(\text{NO}_3)_2$	102	118	139	206							
Manganese(II) oxalate	$\text{MnC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$	0.02	0.024	0.028	0.033							
Manganese(II) sulfate	$\text{MnSO}_4$	52.9	59.7	62.9	62.9	60		53.6		45.6	40.9	35.3
Mercury(I) azide	$\text{Hg}_2(\text{N}_3)_2$			0.02727								
Mercury(I) bromide	$\text{Hg}_2\text{Br}_2$			$1.352 \times 10^{-6}$								
Mercury(I) carbonate	$\text{Hg}_2\text{CO}_3$			$4.351 \times 10^{-7}$								
Mercury(I) chloride	$\text{Hg}_2\text{Cl}_2$			$3.246 \times 10^{-5}$								
Mercury(I) chromate	$\text{Hg}_2\text{CrO}_4$			0.002313								
Mercury(I) cyanide	$\text{Hg}_2(\text{CN})_2$			$2.266 \times 10^{-12}$								
Mercury(I) perchlorate	$\text{Hg}_2(\text{ClO}_4)_2$	282	325	407	455		499		541		580	
Mercury(I) sulfate	$\text{Hg}_2\text{SO}_4$			0.04277								
Mercury(II) acetate	$\text{Hg}(\text{C}_2\text{H}_3\text{O}_2)_2$			25								
Mercury(II) benzoate	$\text{Hg}(\text{C}_7\text{H}_5\text{O}_2)_2 \cdot 2\text{O}$			1.1								
Mercury(II) bromate	$\text{Hg}(\text{BrO}_3)_2 \cdot 2\text{H}_2\text{O}$			0.08								
Mercury(II) bromide	$\text{HgBr}_2$	0.3	0.4	0.56	0.66	0.91		1.68		2.77		4.9

Mercury(II) chlorate	<span>Hg(ClO<sub>3</sub>)<sub>2</sub></span>			25								
Mercury(II) chloride	<span>HgCl<sub>2</sub></span>	3.63	4.82	6.57	8.34	10.2		16.3		30		61.3
Mercury(II) cyanide	<span>Hg(CN)<sub>2</sub></span>			9.3								
Mercury(II) iodate	<span>Hg(IO<sub>3</sub>)<sub>2</sub></span>			0.002372								
Mercury(II) iodide	<span>HgI<sub>2</sub></span>			0.006								
Mercury(II) oxalate	<span>HgC<sub>2</sub>O<sub>4</sub></span>			0.011								
Mercury(II) sulfide	<span>HgS</span>			2.943×10 <sup>−25</sup>								
Mercury(II) thiocyanate	<span>Hg(SCN)<sub>2</sub></span>			0.063								

**N** [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Neodymium(III) acetate	<span>Nd(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub>·H<sub>2</sub>O</span>			26.2								
Neodymium(III) bromate	<span>Nd(BrO<sub>3</sub>)<sub>3</sub></span>	43.9	59.2	75.6	95.2	116						
Neodymium(III) chloride	<span>NdCl<sub>3</sub></span>		96.7	98	99.6	102		105				
Neodymium(III) molybdate	<span>Nd<sub>2</sub>(MoO<sub>4</sub>)<sub>3</sub></span>				0.0019							
Neodymium(III) nitrate	<span>Nd(NO<sub>3</sub>)<sub>3</sub></span>	127	133	142	145	159		211				
Neodymium(III) selenate	<span>Nd<sub>2</sub>(SeO<sub>4</sub>)<sub>3</sub></span>	45.2	44.6	41.8	39.9	39.9		43.9		7	3.3	
Neodymium(III) sulfate	<span>Nd<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></span>	13	9.7	7.1	5.3	4.1		2.8		2.2	1.2	
Nickel(II) acetate	<span>Ni(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></span>											
Nickel(II) bromate	<span>Ni(BrO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O</span>			28								
Nickel(II) bromide	<span>NiBr<sub>2</sub></span>	113	122	131	138	144		153		154		155
Nickel(II) carbonate	<span>NiCO<sub>3</sub></span>			9.643×10 <sup>−4</sup>								
Nickel(II) chlorate	<span>Ni(ClO<sub>3</sub>)<sub>2</sub></span>	111	120	133	155	181		221		308		
Nickel(II) chloride	<span>NiCl<sub>2</sub></span>	53.4	56.3	66.8	70.6	73.2		81.2		86.6		87.6
Nickel(II) fluoride	<span>NiF<sub>2</sub></span>		2.55	2.56				2.56			2.59	
Nickel(II) formate	<span>Ni(HCO<sub>2</sub>)<sub>2</sub>·2 H<sub>2</sub>O</span>		3.15	3.25								
Nickel(II) hydroxide	<span>Ni(OH)<sub>2</sub></span>			0.013								
Nickel(II) iodate	<span>Ni(IO<sub>3</sub>)<sub>2</sub></span>	0.74		0.062	1.43							
Nickel(II) iodide	<span>NiI<sub>2</sub></span>	124	135	148	161	174		184		187	188	
Nickel(II) nitrate	<span>Ni(NO<sub>3</sub>)<sub>2</sub></span>	79.2		94.2	105	119		158		187	188	
Nickel oxalate	<span>NiC<sub>2</sub>O<sub>4</sub>·2H<sub>2</sub>O</span>			0.00118								
Nickel(II) perchlorate	<span>Ni(ClO<sub>4</sub>)<sub>2</sub></span>	105	107	110	113	117						
Nickel(II) pyrophosphate	<span>Ni<sub>2</sub>P<sub>2</sub>O<sub>7</sub></span>			0.001017								
Nickel(II) sulfate	<span>NiSO<sub>4</sub>·6H<sub>2</sub>O</span>			44.4	46.6	49.2		55.6		64.5	70.1	76.7
Nitric oxide	<span>NO</span>			0.0056								
Nitrous oxide	<span>N<sub>2</sub>O</span>			0.112								

**O** [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Oxygen at a partial pressure of 21 kPa	<span>O<sub>2</sub></span>	0.00146	0.00113	0.00091	0.00076	0.00065						
Oxalic acid	<span>H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>·2H<sub>2</sub>O</span>	4.96	8.51	13.3	19.9	30.1		62.1		118	168	

**P** [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100
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												°C
Palladium(II) hydroxide	$\text{Pd(OH)}_2$			$4.106 \times 10^{-10}$								
Palladium(IV) hydroxide	$\text{Pd(OH)}_4$			$5.247 \times 10^{-14}$								
Phenol	$\text{C}_6\text{H}_5\text{OH}$			8.3		miscible						
Platinum(II) hydroxide	$\text{Pt(OH)}_2$			$3.109 \times 10^{-11}$								
Platinum(IV) bromide	$\text{PtBr}_4$			$1.352 \times 10^{-7}$								
Plutonium(III) fluoride	$\text{PuF}_3$			$3.144 \times 10^{-4}$								
Plutonium(IV) fluoride	$\text{PuF}_4$			$3.622 \times 10^{-4}$								
Plutonium(IV) iodate	$\text{Pu(IO}_3)_4$			0.07998								
Polonium(II) sulfide	$\text{PoS}$			$2.378 \times 10^{-14}$								
Potassium acetate	$\text{KC}_2\text{H}_3\text{O}_2$	216	233	256	283	324		350		381	398	
Potassium arsenate	$\text{K}_3\text{AsO}_4$			19								
Potassium azide	$\text{KN}_3$	41.4	46.2	50.8	55.8	61						106
Potassium benzoate	$\text{KC}_7\text{H}_5\text{O}_2$		65.8	70.7	76.7	82.1						
Potassium bromate	$\text{KBrO}_3$	3.09	4.72	6.91	9.64	13.1		22.7		34.1		49.9
Potassium bromide	$\text{KBr}$	53.6	59.5	65.3	70.7	75.4		85.5		94.9	99.2	104
Potassium hexabromoplatinate	$\text{K}_2\text{PtBr}_6$			1.89								
Potassium carbonate	$\text{K}_2\text{CO}_3$	105	109	111	114	117	121.2	127		140	148	156
Potassium chlorate	$\text{KClO}_3$	3.3	5.2	7.3	10.1	13.9		23.8		37.5	46	56.3
Potassium chloride	$\text{KCl}$	28	31.2	34.2	37.2	40.1	42.6	45.8		51.3	53.9	56.3
Potassium chromate	$\text{K}_2\text{CrO}_4$	56.3	60	63.7	66.7	67.8		70.1			74.5	
Potassium cyanide	$\text{KCN}$			50								
Potassium dichromate	$\text{K}_2\text{Cr}_2\text{O}_7$	4.7	7	12.3	18.1	26.3	34	45.6		73		
Potassium dihydrogen arsenate	$\text{KH}_2\text{AsO}_4$			19								
Potassium dihydrogen phosphate	$\text{KH}_2\text{PO}_4$	14.8	18.3	22.6	28	35.5	41	50.2		70.4	83.5	
Potassium ferricyanide	$\text{K}_3\text{Fe(CN)}_6$	30.2	38	46	53	59.3		70				91
Potassium ferrocyanide	$\text{K}_4\text{Fe(CN)}_6$	14.3	21.1	28.2	35.1	41.4		54.8		66.9	71.5	74.2
Potassium fluoride	$\text{KF}$	44.7	53.5	94.9	108	138		142		150		
Potassium formate	$\text{KHCO}_2$	32.8	313	337	361	398		471		580	658	
Potassium hydrogen carbonate	$\text{KHCO}_3$	22.5	27.4	33.7	39.9	47.5		65.6				
Potassium hydrogen phosphate	$\text{K}_2\text{HPO}_4$			150								
Potassium hydrogen sulfate	$\text{KHSO}_4$	36.2		48.6	54.3	61		76.4		96.1		122
Potassium hydrogen tartrate	$\text{KHC}_4\text{H}_4\text{O}_6$			0.6								6,2
Potassium hydroxide	$\text{KOH}$	95.7	103	112	126	134		154				178
Potassium iodate	$\text{KIO}_3$	4.6	6.27	8.08	10.3	12.6	14	18.3		24.8		32.3
Potassium iodide	$\text{KI}$	128	136	144	153	162		176		192	198	206
Potassium												

metabisulfite	$K_2S_2O_5$				45							
Potassium nitrate	$KNO_3$	13	22	33	48	65	84	106	132	167	199.5	240
Potassium nitrite	$KNO_2$	279	292	306	320	329		348		376	390	410
Potassium oxalate	$K_2C_2O_4$	25.5	31.9	36.4	39.9	43.8		53.2		63.6	69.2	75.3
Potassium perchlorate	$KClO_4$	0.76	1.06	1.68	2.56	3.73		7.3		13.4	17.7	22.3
Potassium periodate	$KIO_4$	0.17	0.28	0.42	0.65	1		2.1		4.4	5.9	
Potassium permanganate	$KMnO_4$	2.83	4.31	6.34	9.03	12.6	16.9	22.1				
Potassium persulfate	$K_2S_2O_8$			4.7								
Potassium phosphate	$K_3PO_4$		81.5	92.3	108	133						
Potassium selenate	$K_2SeO_4$	107	109	111	113	115		119		121		122
Potassium sulfate	$K_2SO_4$	7.4	9.3	11.1	13	14.8		18.2		21.4	22.9	24.1
Potassium tetraphenylborate	$KBC_{24}H_{20}$			$1.8 \times 10^{-5}$								
Potassium thiocyanate	$KSCN$	177	198	224	255	289		372		492	571	675
Potassium thiosulfate	$K_2S_2O_3$	96		155	175	205		238		293	312	
Potassium tungstate	$K_2WO_4$			51.5								
Praseodymium(III) acetate	$Pr(C_2H_3O_2)_3 \cdot H_2O$			32								
Praseodymium(III) bromate	$Pr(BrO_3)_3$	55.9	73	91.8	114	144						
Praseodymium(III) chloride	$PrCl_3$			104								
Praseodymium(III) molybdate	$Pr_2(MoO_4)_3$			0.0015								
Praseodymium(III) nitrate	$Pr(NO_3)_3$			112	162	178						
Praseodymium(III) sulfate	$Pr_2(SO_4)_3$	19.8	15.6	12.6	9.89	2.56		5.04		3.5	1.1	0.91

## R [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Radium chloride	$RaCl_2$			19.6								
Radium iodate	$Ra(IO_3)_2$			0.04								
Radium nitrate	$Ra(NO_3)_2$			12								
Radium sulfate	$RaSO_4$			$2.1 \times 10^{-4}$								
Raffinose	$C_{18}H_{32}O_{16} \cdot 5H_2O$			14								
Rubidium acetate	$RbC_2H_3O_2$					86						
Rubidium bromate	$RbBrO_3$				3.6	5.1						
Rubidium bromide	$RbBr$	90	99	108	119	132		158				
Rubidium chlorate	$RbClO_3$	2.1	3.1	5.4	8	11.6		22		38	49	63
Rubidium chloride	$RbCl$	77	84	91	98	104		115		127	133	143
Rubidium chromate	$Rb_2CrO_4$	62	67.5	73.6	78.9	85.6		95.7				
Rubidium dichromate	$Rb_2Cr_2O_7$			5.9	10	15.2		32.3				
Rubidium fluoride	$RbF$			300								
Rubidium fluorosilicate	$Rb_2SiF_6$			0.157								
Rubidium formate	$RbHCO_2$		443	554	614	694		900				
Rubidium hydrogen carbonate	$RbHCO_3$			110								

Rubidium hydroxide	RbOH			180								
Rubidium iodate	RbIO <sub>3</sub>			1.96								
Rubidium iodide	RbI			144								
Rubidium nitrate	RbNO <sub>3</sub>	19.5	33	52.9	81.2	117		200		310	374	452
Rubidium perchlorate	RbClO <sub>4</sub>	1.09	1.19	1.55	2.2	3.26		6.27		11	15.5	22
Rubidium periodate	RbIO <sub>4</sub>			0.648								
Rubidium selenate	Rb <sub>2</sub> SeO <sub>4</sub>			159								
Rubidium sulfate	Rb <sub>2</sub> SO <sub>4</sub>	37.5	42.6	48.1	53.6	58.5		67.5		75.1	78.6	81.8

## S [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Samarium acetate	Sm(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>3</sub> ·3H <sub>2</sub> O			15								
Samarium bromate	Sm(BrO <sub>3</sub> ) <sub>3</sub>	34.2	47.6	62.5	79	98						
Samarium chloride	SmCl <sub>3</sub>		92.4	93.4	94.6	96.9						
Samarium sulfate	Sm <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> ·8H <sub>2</sub> O			2.7	3.1							
Scandium oxalate	Sc <sub>2</sub> (C <sub>2</sub> O <sub>4</sub> ) <sub>3</sub> ·6H <sub>2</sub> O			0.006								
Scandium sulfate	Sc <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> ·5H <sub>2</sub> O			54.6								
Silicon dioxide	SiO <sub>2</sub>			0.012								
Silver acetate	AgC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	0.73	0.89	1.05	1.23	1.43		1.93		2.59		
Silver azide	AgN <sub>3</sub>			7.931×10 <sup>-4</sup>								
Silver bromate	AgBrO <sub>3</sub>		0.11	0.16	0.23	0.32		0.57		0.94	1.33	
Silver bromide	AgBr			1.328×10 <sup>-5</sup>								
Silver carbonate	Ag <sub>2</sub> CO <sub>3</sub>			0.003489								
Silver chlorate	AgClO <sub>3</sub>		10.4	15.3	20.9	26.8						
Silver chloride	AgCl			1.923×10 <sup>-4</sup>			5.2×10 <sup>-5</sup>					
Silver chlorite	AgClO <sub>2</sub>			0.248								
Silver chromate	Ag <sub>2</sub> CrO <sub>4</sub>			0.002157								
Silver cyanide	AgCN			1.467×10 <sup>-7</sup>								
Silver dichromate	Ag <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub>			0.159								
Silver fluoride	AgF	85.9	120	172	190	203						
Silver nitrate	AgNO <sub>3</sub>	122	167	216	265	311		440		585	652	733
Silver oxalate	Ag <sub>2</sub> C <sub>2</sub> O <sub>4</sub>			0.00327								
Silver oxide	Ag <sub>2</sub> O			0.0012								
Silver perchlorate	AgClO <sub>4</sub>	455	484	525	594	635						793
Silver permanganate	AgMnO <sub>4</sub>			0.9								
Silver sulfate	Ag <sub>2</sub> SO <sub>4</sub>	0.57	0.7	0.8	0.89	0.98		1.15		1.3	1.36	1.41
Silver vanadate	AgVO <sub>3</sub>			0.01462								
Sodium acetate	NaC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	36.2	40.8	46.4	54.6	65.6		139		153	161	170
Sodium azide	NaN <sub>3</sub>	38.9	39.9	40.8								
Sodium benzoate	NaC <sub>7</sub> H <sub>5</sub> O <sub>2</sub>			66								
Sodium borohydride	NaBH <sub>4</sub>	25		55		88.5						
Sodium bromate	NaBrO <sub>3</sub>	24.2	30.3	36.4	42.6	48.8		62.6		75.7		90.8
Sodium bromide	NaBr	80.2	85.2	90.8	98.4	107		118		120	121	121
Sodium carbonate	Na <sub>2</sub> CO <sub>3</sub>	7	12.5	21.5	39.7	49		46		43.9	43.9	45.5
Sodium chlorate	NaClO <sub>3</sub>	79.6	87.6	95.9	105	115		137		167	184	204
Sodium chloride	NaCl	35.65	35.72	35.89	36.09	36.37	36.69	37.04	37.46	37.93	38.47	38.99

Sodium chromate	$\text{Na}_2\text{CrO}_4$	31.7	50.1	84	88	96		115		125		126
Sodium cyanide	$\text{NaCN}$	40.8	48.1	58.7	71.2	dec						
Sodium dichromate	$\text{Na}_2\text{Cr}_2\text{O}_7$	163	172	183	198	215		269		376	405	415
Monosodium phosphate	$\text{NaH}_2\text{PO}_4$	56.5	69.8	86.9	107	133		172		211	234	
Sodium fluoride	$\text{NaF}$	3.66		4.06	4.22	4.4		4.68		4.89		5.08
Sodium formate	$\text{HCOONa}$	43.9	62.5	81.2	102	108		122		138	147	160
Sodium hydrogen carbonate	$\text{NaHCO}_3$	7	8.1	9.6	11.1	12.7		16				
Sodium hydroxide	$\text{NaOH}$		98	109	119	129		174				
Sodium iodate	$\text{NaIO}_3$	2.48	4.59	8.08	10.7	13.3		19.8		26.6	29.5	33
Sodium iodide	$\text{NaI}$	159	167	178	191	205		257		295		302
Sodium metabisulfite	$\text{Na}_2\text{S}_2\text{O}_5$	45.1		65.3						88.7		96.3
Sodium metaborate	$\text{NaBO}_2$	16.4	20.8	25.4	31.4	40.4		63.9		84.5		125.2
Sodium molybdate	$\text{Na}_2\text{MoO}_4$	44.1	64.7	65.3	66.9	68.6		71.8				
Sodium nitrate	$\text{NaNO}_3$	73	80.8	87.6	94.9	102		122		148		180
Sodium nitrite	$\text{NaNO}_2$	71.2	75.1	80.8	87.6	94.9		111		133		160
Sodium oxalate	$\text{Na}_2\text{C}_2\text{O}_4$	2.69	3.05	3.41	3.81	4.18		4.93		5.71		6.5
Sodium perchlorate	$\text{NaClO}_4$	167	183	201	222	245		288		306		329
Sodium periodate	$\text{NaIO}_4$	1.83	5.6	10.3	19.9	30.4						
Sodium permanganate	$\text{NaMnO}_4$			90								
Sodium phosphate	$\text{Na}_3\text{PO}_4$	4.5	8.2	12.1	16.3	20.2		20.9		60	68.1	77
Sodium pyrophosphate	$\text{Na}_4\text{P}_2\text{O}_7$	2.26										
Sodium selenate	$\text{Na}_2\text{SeO}_4$	13.3	25.2	26.9	77	81.8		78.6		74.8	73	72.7
Sodium sulfate	$\text{Na}_2\text{SO}_4$	4.9	9.1	19.5	40.8	48.8		45.3		43.7	42.7	42.5
Sodium sulfite	$\text{Na}_2\text{SO}_3$				27.0							
Sodium tetraborate decahydrate	$\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$	2	2.3	2.5	4	6	10	15				
Sodium tetraborate pentahydrate	$\text{Na}_2\text{B}_4\text{O}_7 \cdot 5\text{H}_2\text{O}$								20	23	28	35
Sodium tetraborate tetrahydrate	$\text{Na}_2\text{B}_4\text{O}_7 \cdot 4\text{H}_2\text{O}$								17	20	23	28
Sodium tetraphenylborate	$\text{NaB}(\text{C}_6\text{H}_5)_4$			47								
Sodium thiosulfate	$\text{Na}_2\text{S}_2\text{O}_3$	71.5		73		77.6				90.8		97.2
Strontium acetate	$\text{Sr}(\text{C}_2\text{H}_3\text{O}_2)_2$	37	42.9	41.1	39.5	38.3		36.8		36.1	36.2	36.4
Strontium bromate	$\text{Sr}(\text{BrO}_3)_2 \cdot \text{H}_2\text{O}$			30.9								41
Strontium bromide	$\text{SrBr}_2$	85.2	93.4	102	112	123		150		182		223
Strontium carbonate	$\text{SrCO}_3$			0.0011								0.065
Strontium chlorate	$\text{SrClO}_3$			175								
Strontium chloride	$\text{SrCl}_2$	43.5	47.7	52.9	58.7	65.3		81.8		90.5		101
Strontium chromate	$\text{SrCrO}_4$			0.085	0.090							

Strontium fluoride	$\text{SrF}_2$			$1.2 \times 10^{-4}$								
Strontium formate	$\text{Sr}(\text{HCO}_2)_2$	9.1	10.6	12.7	15.2	17.8		25		31.9	32.9	34.4
Strontium hydroxide	$\text{Sr}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$	0.91	1.25	1.77	2.64	3.95		8.42		20.2	44.5	91.2
Strontium iodate	$\text{Sr}(\text{IO}_3)_2$			0.19								0.35
Strontium iodide	$\text{SrI}_2$	165		178		192		218		270	365	383
Strontium molybdate	$\text{SrMoO}_4$			0.01107								
Strontium nitrate	$\text{Sr}(\text{NO}_3)_2$	39.5	52.9	69.5	88.7	89.4		93.4		96.9	98.4	
Strontium selenate	$\text{SrSeO}_4$			0.656								
Strontium sulfate	$\text{SrSO}_4$	0.0113	0.0129	0.0132	0.0138	0.0141		0.0131		0.0116	0.0115	
Strontium thiosulfate	$\text{SrS}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$		2.5									
Strontium tungstate	$\text{SrWO}_4$			$3.957 \times 10^{-4}$								
Sucrose	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$	181.9	190.6	201.9	216.7	235.6	259.6	288.8	323.7	365.1	414.9	476.0
Sulfur dioxide	$\text{SO}_2$			9.4								

## T [edit]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Terbium bromate	$\text{Tb}(\text{BrO}_3)_3 \cdot 9\text{H}_2\text{O}$	66.4	89.7	117	152	198						
Terbium sulfate	$\text{Tb}_2(\text{SO}_4)_3 \cdot 8\text{H}_2\text{O}$			3.56								
Thallium(I) azide	$\text{TlN}_3$	0.171	0.236	0.364								
Thallium(I) bromate	$\text{TlBrO}_3$			0.306								
Thallium(I) bromide	$\text{TlBr}$	0.022	0.032	0.048	0.068	0.097		0.117				
Thallium(I) carbonate	$\text{Tl}_2\text{CO}_3$			5.3								
Thallium(I) chlorate	$\text{TlClO}_3$	2		3.92		12.7				36.6		57.3
Thallium(I) cyanide	$\text{TlCN}$			16.8								
Thallium(I) fluoride	$\text{TlF}$			78								
Thallium(I) hydrogen carbonate	$\text{TlHCO}_3$			500								
Thallium(I) hydroxide	$\text{TlOH}$	25.4	29.6	35	40.4	49.4		73.3		106	126	150
Thallium(I) iodate	$\text{TlIO}_3$			0.06678								
Thallium(I) iodide	$\text{TlI}$	0.002		0.006		0.015		0.035		0.07		0.12
Thallium(I) nitrate	$\text{TlNO}_3$	3.9	6.22	9.55	14.3	21		46.1		110	200	414
Thallium(I) oxalate	$\text{Tl}_2\text{C}_2\text{O}_4$			1.83								
Thallium(I) perchlorate	$\text{TlClO}_4$	6	8.04	13.1	19.7	28.3		50.8		81.5		
Thallium(I) phosphate	$\text{Tl}_3\text{PO}_4$			0.15								
Thallium(I) pyrophosphate	$\text{Tl}_4\text{P}_2\text{O}_7$			40								
Thallium(I) selenate	$\text{Tl}_2\text{SeO}_4$		2.17	2.8						8.5		10.8
Thallium(I) sulfate	$\text{Tl}_2\text{SO}_4$	2.73	3.7	4.87	6.16	7.53		11		14.6	16.5	18.4



Thallium(I) vanadate	$\text{TlVO}_3$			0.87									
Thorium(IV) fluoride	$\text{ThF}_4 \cdot 4\text{H}_2\text{O}$			0.914									
Thorium(IV) iodate	$\text{Th}(\text{IO}_3)_4$			0.03691									
Thorium(IV) nitrate	$\text{Th}(\text{NO}_3)_4$	186	187	191									
Thorium(IV) selenate	$\text{Th}(\text{SeO}_4)_2 \cdot 9\text{H}_2\text{O}$	0.65											
Thorium(IV) sulfate	$\text{Th}(\text{SO}_4)_2 \cdot 9\text{H}_2\text{O}$	0.74	0.99	1.38	1.99	3							
Tin(II) bromide	$\text{SnBr}_2$	85											
Tin(II) chloride	$\text{SnCl}_2$	84											
Tin(II) fluoride	$\text{SnF}_2$			30									
Tin(II) iodide	$\text{SnI}_2$			0.99	1.17	1.42		2.11		3.04	3.58	4.2	
Tin(II) sulfate	$\text{SnSO}_4$			18.9									
Trehalose	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$			68.9									

## U [edit]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Uranyl acetate	$\text{UO}_2(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 2\text{H}_2\text{O}$			7.69								
Uranyl chloride	$\text{UO}_2\text{Cl}_2$			320								
Uranyl formate	$\text{UO}_2(\text{HCO}_2)_2 \cdot \text{H}_2\text{O}$			7.2								
Uranyl iodate	$\text{UO}_2(\text{IO}_3)_2 \cdot \text{H}_2\text{O}$			0.124								
Uranyl nitrate	$\text{UO}_2(\text{NO}_3)_2$	98	107	122	141	167		317		388	426	474
Uranyl oxalate	$\text{UO}_2\text{C}_2\text{O}_4$		0.45	0.5	0.61	0.8		1.22		1.94		3.16
Uranyl sulfate	$\text{UO}_2\text{SO}_4 \cdot 3\text{H}_2\text{O}$			21								
Urea	$\text{CO}(\text{NH}_2)_2$			108		167		251		400		733

## V [edit]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Vanadium(V) oxide	$\text{V}_2\text{O}_5$			0.8								

## X [edit]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Xenon	$\text{Xe}$	24.1 ml		11.9 ml <sup>25</sup>			8.4 ml			7.12 ml		
Xylose	$\text{C}_5\text{H}_{10}\text{O}_5$			117								

## Y [edit]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Ytterbium(III) sulfate	$\text{Yb}_2(\text{SO}_4)_3$	44.2	37.5		22.2	17.2		10.4		6.4	5.8	4.7
Yttrium(III) acetate	$\text{Y}(\text{C}_2\text{H}_3\text{O}_2)_3 \cdot 4\text{H}_2\text{O}$			9.03								
Yttrium(III) bromate	$\text{Y}(\text{BrO}_3)_3 \cdot 9\text{H}_2\text{O}$			168								
Yttrium(III) bromide	$\text{YBr}_3$	63.9		75.1		87.3		101		116	123	
Yttrium(III) chloride	$\text{YCl}_3$	77.3	78.1	78.8	79.6	80.8						
Yttrium(III) fluoride	$\text{YF}_3$			0.005769								

Yttrium(III) nitrate	$Y(NO_3)_3$	93.1	106	123	143	163		200				
Yttrium(III) sulfate	$Y_2(SO_4)_3$	8.05	7.67	7.3	6.78	6.09		4.44		2.89	2.2	

**Z** [[edit](#)]

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Zinc acetate	$Zn(C_2H_3O_2)_2$			30								
Zinc bromide	$ZnBr_2$	389		446	528	591		618		645		672
Zinc carbonate	$ZnCO_3$			$4.692 \times 10^{-5}$								
Zinc chlorate	$Zn(ClO_3)_2$	145	152	209	223							
Zinc chloride	$ZnCl_2$	342	353	395	437	452		488		541		614
Zinc cyanide	$Zn(CN)_2$			0.058								
Zinc fluoride	$ZnF_2$			1.6								
Zinc formate	$Zn(HCO_2)_2$	3.7	4.3	6.1	7.4		11.8		21.2	28.8	38	
Zinc iodate	$Zn(IO_3)_2 \cdot 2H_2O$			0.07749								
Zinc iodide	$ZnI_2$	430		432		445		467		490		510
Zinc nitrate	$Zn(NO_3)_2$	98			138	211						
Zinc oxalate	$ZnC_2O_4 \cdot 2H_2O$			$1.38 \times 10^{-9}$								
Zinc permanganate	$Zn(MnO_4)_2$			33.3								
Zinc sulfate	$ZnSO_4$	41.6	47.2	53.8	61.3	70.5		75.4		71.1		60.5
Zinc sulfite	$ZnSO_3 \cdot 2H_2O$			0.16								
Zinc tartrate	$ZnC_4H_4O_6$			0.022	0.041	0.06		0.104		0.59		
Zirconium fluoride	$ZrF_4$			1.32								
Zirconium sulfate	$Zr(SO_4)_2 \cdot 4H_2O$			52.5								

External links [[edit](#)]

- [Solubility Database](#) - International Union of Pure and Applied Chemistry / National Institute of Standards and Technology

References [[edit](#)]

- Chemicalc v4.0 - a software that includes data on solubility
- [\[1\]](#) [Learning, Food resources](#)
- [\[2\]](#) [Kaye and Laby Online](#)
- [\[3\]](#) [Chemfinder.com](#)

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